

Oxidation-Reduction Worksheet

For each reaction below, identify the atom oxidized, the atom reduced, the oxidizing agent, the reducing agent, the oxidation half reaction, the reduction half reaction, and then balance the equation by the method of oxidation-reduction showing all electrons transfers.

1. $\overset{0}{\text{Mg}} + 2\overset{+1}{\text{H}}\overset{-1}{\text{Cl}} \rightarrow \overset{+2}{\text{Mg}}\overset{-2}{\text{Cl}}_2 + \overset{0}{\text{H}}_2$ $\overset{0}{\text{Mg}} \xrightarrow{-2e^-} \overset{+2}{\text{Mg}}$ $2\overset{+1}{\text{H}} \xrightarrow{+2e^-} \overset{0}{\text{H}}_2$
2. $2\overset{0}{\text{Fe}} + 3\overset{+3}{\text{V}}_2\overset{-2}{\text{O}}_3 \rightarrow \overset{+3}{\text{Fe}}_2\overset{-2}{\text{O}}_3 + 6\overset{+2}{\text{V}}\overset{-2}{\text{O}}$ $2\overset{0}{\text{Fe}} \xrightarrow{-6e^-} \overset{+3}{\text{Fe}}_2$ $3\overset{+3}{\text{V}}_2 \xrightarrow{+6e^-} 6\overset{+2}{\text{V}}$
3. $\overset{+7}{\text{K}}\overset{+4}{\text{Mn}}\overset{-2}{\text{O}}_4 + \overset{+5}{\text{K}}\overset{-2}{\text{NO}}_2 + 2\overset{+2}{\text{H}}_2\overset{-2}{\text{SO}}_4 \rightarrow \overset{+2}{\text{Mn}}\overset{-2}{\text{SO}}_4 + 2\overset{+2}{\text{H}}_2\overset{-2}{\text{O}} + \overset{+5}{\text{K}}\overset{-2}{\text{NO}}_3 + \overset{+6}{\text{K}}_2\overset{-2}{\text{SO}}_4$
4. $\text{K}_2\text{Cr}_2\text{O}_7 + \text{SnCl}_2 + \text{HCl} \rightarrow \text{CrCl}_3 + \text{SnCl}_4 + \text{KCl} + \text{H}_2\text{O}$
5. $\text{KMnO}_4 + \text{NaCl} + \text{H}_2\text{SO}_4 \rightarrow \text{Cl}_2 + \text{K}_2\text{SO}_4 + \text{MnSO}_4 + \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$
6. $\text{K}_2\text{Cr}_2\text{O}_7 + \text{H}_2\text{O} + \text{S} \rightarrow \text{SO}_2 + \text{KOH} + \text{Cr}_2\text{O}_3$
7. $\text{KClO}_3 + \text{C}_{12}\text{H}_{22}\text{O}_{11} \rightarrow \text{KCl} + \text{H}_2\text{O} + \text{CO}_2$
8. $\text{H}_2\text{C}_2\text{O}_4 + \text{K}_2\text{MnO}_4 \rightarrow \text{CO}_2 + \text{K}_2\text{O} + \text{Mn}_2\text{O}_3 + \text{H}_2\text{O}$
9. $\text{Mn}(\text{NO}_3)_2 + \text{NaBiO}_3 + \text{HNO}_3 \rightarrow \text{HMnO}_4 + \text{Bi}(\text{NO}_3)_3 + \text{NaNO}_3 + \text{H}_2\text{O}$
10. $\text{H}_2\text{C}_2\text{O}_4 + \text{KMnO}_4 \rightarrow \text{CO}_2 + \text{K}_2\text{O} + \text{Mn}_2\text{O}_3 + \text{H}_2\text{O}$

Analyzing Oxidation Reduction Reactions Answer Key

**Dr. Santosh Karajgi, Prof. (Dr.) Anil
Ahuja, Dr. Ramesh Jayaprakash, Dr.
Kella Alekhya, Dr. D. Sudhahar**

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Chemical Engineering Design and Analysis T. Michael Duncan, Jeffrey A. Reimer, 2019-01-24 The go to guide to learn the principles and practices of design and analysis in chemical engineering

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Analysis of the Oxidation of Chalcopyrite, Chalcocite, Galena, Pyrrhotite, Marcasite, and Arsenopyrite George W. Reimers, Kenneth E. Hjelmstad, 1987

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Publications of the National Bureau of Standards United States. National Bureau of Standards,1985

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A Textbook of Pharmaceutical Analysis Dr. Santosh Karajgi, Prof. (Dr.) Anil Ahuja, Dr. Ramesh Jayaprakash, Dr. Kella Alekhya, Dr. D. Sudhahar,2024-01-31

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