

Lab objective:

Materials:

- | | |
|------------------------------------|-----------------------|
| - Vinegar (5% acetic acid) | - Graduated cylinders |
| - Baking soda (sodium bicarbonate) | - Ziploc bag |
| - Electronic scale | - Calculator |

Procedure:

1. Balance the chemical reaction of baking soda and vinegar:



2. Get a Ziploc bag
3. Record:
temperature: _____ Pressure: _____ in the room.
4. Determine the volume of your Ziploc bag in liters: _____
5. Calculate the mass of baking soda (sodium bicarbonate) required to completely fill your airbag with CO_2 .
6. Calculate the volume vinegar (5% acetic acid) required to completely fill your airbag. Vinegar has an acetic acid, $\text{HC}_2\text{H}_3\text{O}_2$, concentration of 0.833 mol / L.

Airbag Stoichiometry Lab Work And Answers

JR Anderson



Airbag Stoichiometry Lab Work And Answers:

Improving Student Comprehension of Stoichiometric Concepts Connie Lynn Bannick Kemner,2007 Improving Student Comprehension in Chemistry Laboratories Tracy Lynn Haroff,2006 O-level Chemistry Challenging Learn-By-Example (Yellowreef) Thomas Bond,Chris Hughes,2013-11-19 covers latest MOE syllabus comprehensive examples and solutions for quick revision helps students to familiarise with various exam question types complete edition and concise edition eBooks available **Annual Index/abstracts of SAE Technical Papers** ,1999 **Energy Conservation Update** ,1979 **A Stoichiometry Unit** David Callaghan,2004

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