

Algebra 1 Fall Semester Final Exam Test Bank

Indicate the answer choice that best completes the statement or answers the question.

1. Write an algebraic expression for *the sum of a number and 8*.
a. $8x$ b. $x - 8$ c. $x + 8$ d. $x \div 8$
2. Write an algebraic expression for *27 decreased by a number*.
a. $27 + m$ b. $27 - m$ c. $m - 27$ d. $\frac{27}{m}$
3. Write an algebraic expression for *3 times x squared minus 4 times x* .
a. $3(2x) - 4x$ b. $4 - 3x$ c. $3x^2 - 4x$ d. $3(x^2 - 4x)$
4. Write an algebraic expression for *nine times the square of a number*.
a. $9 + x^2$ b. $9 - x^2$ c. $9x^2$ d. $x^2 - 9$
5. Write a verbal expression for $19a$.
a. the sum of 19 and a number b. the difference of 19 and a number
c. the quotient of 19 and a number d. the product of 19 and a number
6. Write a verbal expression for $2n + 7$.
a. the product of 2, n , and 7 b. 7 more than twice a number
c. 7 less than a number times 2 d. 7 more than n and 2
7. Write a verbal expression for $3n - 8$.
a. the product of 3, n , and 8 b. 8 less than the product of 3 and n
c. 3 times n less than 8 d. n minus 8 times 3
8. Evaluate $6 + 2 \cdot 3 - 1$.
a. 23 b. 10 c. 16 d. 11
9. Evaluate $4 + 5 \cdot 7 - 1$.
a. 139 b. 15 c. 34 d. 38

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hydrogen

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web jan 10 2023 introduction in this lab we will apply the concepts of reaction rate order of reaction and catalyst our example reaction is the iodination of acetone catalyzed by hydrochloric acid the reaction is zero order in iodine you will figure out the order in acetone and the order in hcl and the rate constant

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web video answer there can a 6 0 liter solution in sodium hydroxide ensure is been reduced go 400 liters so we have to find the moles in the solution the first th

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web the iodination of acetone is convenient because the color of the iodine reactant is readily observable and the change in concentration of 4 the iodine can be easily monitored by our spectrophotometer and the reaction proceeds quickly enough for multiple trials but slowly enough to obtain measurements

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web i the iodination of acetone the rate of this reaction is found to depend on the concentration of hydrogen ion in the solution as well as presumably on the concentrations of the two reactants

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