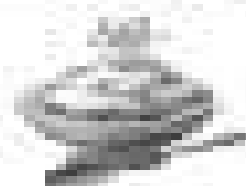


Atomic Probing Structure of the Atom

Name _____
Date _____



Purpose:

To determine the structure of the atom by observing the deflection of alpha particles.

Part I: Finding Evidence for Nuclear Structure

Alpha particles are positively charged particles with a mass of approximately 4 atomic mass units. They are emitted from the nucleus of a radioactive substance. In the Rutherford experiment, alpha particles are directed at a thin foil of gold. The deflection of the particles is observed by a fluorescent screen.

Observation: Most alpha particles pass straight through the foil, but a small number are deflected at large angles.

This observation suggests that the atom is mostly empty space, with a small, dense, positively charged nucleus.

Conclusion: The results of the Rutherford experiment suggest that the atom is mostly empty space, with a small, dense, positively charged nucleus. This is the nuclear model of the atom.

The Rutherford experiment is a classic example of the scientific method, where a hypothesis is tested by observation and experiment.

Observation: The results of the Rutherford experiment suggest that the atom is mostly empty space, with a small, dense, positively charged nucleus.

This observation suggests that the atom is mostly empty space, with a small, dense, positively charged nucleus.

Answers Lesson 11 Atomic Pudding

Didier Musso



Answers Lesson 11 Atomic Pudding:

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Unveiling the Energy of Verbal Artistry: An Mental Sojourn through **Answers Lesson 11 Atomic Pudding**

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