

100 gm solution contains 20 gm  $H_2SO_4$   
 $\frac{100}{d}$  ml solution contains 20 gm of  $H_2SO_4$   
 or  
 1000 ml solution contains  $3.6 \times 98$  gm of  $H_2SO_4$   
 $\therefore 3.6 \times 98 = \frac{20 \times d}{100} = 1000$

$d = 1.23$   
 Hence, (3) is correct.

118. The first and second dissociation constants of an acid  $H_2A$  are  $1.0 \times 10^{-5}$  and  $5.0 \times 10^{-10}$  respectively. The overall dissociation constant of the acid will be

(1)  $5.0 \times 10^{-5}$  (2)  $5.0 \times 10^{-10}$   
 (3)  $5.0 \times 10^{-15}$  (4)  $5.0 \times 10^5$

Ans. (3)



For the reaction



$$K = \frac{[H^+]^2[A^{2-}]}{[H_2A]} = K_1 \times K_2$$

$$= 1 \times 10^{-5} \times 5 \times 10^{-10}$$

$$= 5 \times 10^{-15}$$

Hence, (3) is correct.

119. A mixture of ethyl alcohol and propyl alcohol has a vapour pressure of 290 mm at 300 K. The vapour pressure of propyl alcohol is 200 mm. If the mole fraction of ethyl alcohol is 0.6, its vapour pressure (in mm) at the same temperature will be

(1) 350 (2) 300  
 (3) 700 (4) 360

Ans. (1)

Sol. Let the vapour pressure of pure ethyl alcohol be P.

According to Raoult's law  
 $290 = 200 \times 0.4 + P \times 0.6$

$$P = \frac{290 - 80}{0.6} = 350 \text{ mm Hg}$$

Hence, (1) is correct.

120. In conversion of lime-stone to lime,



the values of  $\Delta H^\circ$  and  $\Delta S^\circ$  are  $+178.1 \text{ kJ mol}^{-1}$  and  $160.2 \text{ J/K}$  respectively at 298 K and 1 bar. Assuming that  $\Delta H^\circ$  do not change with temperature, temperature above which conversion of limestone to lime will be spontaneous is

(1) 1508 K (2) 1200  
 (3) 845 K (4) 1118 K

Ans. (4)

Sol. We know,  $\Delta G = \Delta H - T\Delta S$

So, let's find the equilibrium temperature, i.e. at which

$$\Delta G = 0$$

$$\Delta H = T\Delta S$$

$$T = \frac{178.1 \times 1000}{160.2}$$

$$= 1118 \text{ K}$$

So, at temperature above this, the reaction will become spontaneous.

Hence, (4) is correct answer.

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