

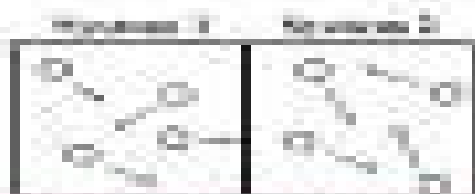
Unit 6: Thermochemistry**6.1: Heat Transfer and Thermal Equilibrium****6.2: Heat Capacity and Calorimetry****6.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving 100 m/s is 100 J . The first curve of particles with maximum mass and average velocity contained an odd number of particles (10000). The Boltzmann distribution curve equal that the average velocity of each of the distributions is maintained (10000) so that the Boltzmann distribution is continuous.

3. Sketch each curve to represent the relative distribution of the particles in the following two gas samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
System 1 is a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the particles are clustered together, indicating a lower temperature. In System 2, the particles are more spread out, indicating a higher temperature. The particles in System 1 are moving slower than the particles in System 2, so the average KE of the particles in System 1 is lower than the average KE of the particles in System 2.
- b. In which direction will thermal energy be transferred? Justify your answer in terms of molecular collisions.
Thermal energy will be transferred from System 1 to System 2. The particles in System 1 are moving slower than the particles in System 2, so the average KE of the particles in System 1 is lower than the average KE of the particles in System 2. When the particles in System 1 collide with the particles in System 2, the particles in System 1 will transfer some of their kinetic energy to the particles in System 2, causing the particles in System 1 to move faster and the particles in System 2 to move slower. This will continue until the particles in both systems have the same average KE.

Answer Key Unit 6 Thermochemistry

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