

8. Buffer, passing question 10, is shown within the Earth's crust. Buffer is composed of the components of those of existing rocks.

A sample of Buffer from a sample was analysed to give the following composition of samples.

sample	composition (%)
1%	0.001
1%	0.01
1%	0.02

- (a) Define the term relative atomic mass.

The relative atomic mass of an element is the mass of one atom of the element divided by the mass of one atom of ^{12}C .

- (b) Calculate the relative atomic mass of the sample of buffer.

Give your answer to three decimal places.

$$\left(\frac{1}{100} \times 12 \right) + \left(\frac{1}{100} \times 16 \right) + \left(\frac{1}{100} \times 16 \right) = \frac{1}{100} \times 100$$

100.00

$$\text{answer} = \frac{100.00}{100} = 1.000$$

- (c) Some studies, as early as 1950, suggested that elements were made up of tiny particles called atoms which could be divided into protons, electrons and neutrons.

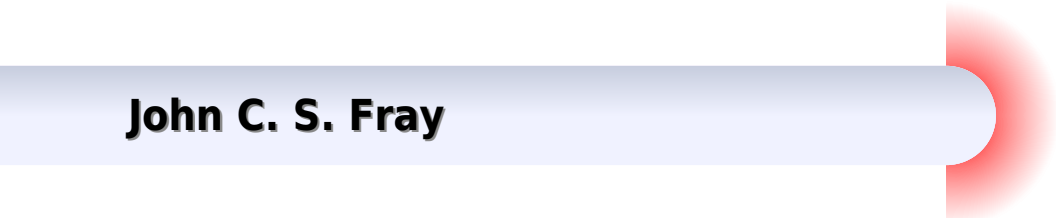
Complete the table to show the number of subatomic particles in the ^{12}C atom and $^{12}\text{C}^{2+}$ ion.

	protons	neutrons	electrons
^{12}C	6	6	6
$^{12}\text{C}^{2+}$	6	6	4



Answer To F321 Chemistry May Series

John C. S. Fray



Answer To F321 Chemistry May Series:

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