

OpenQuest 13

Orbital Configuration

Name _____

Date _____

Class _____

Information: Energy of Sublevels

Not all orbitals have the same energy of energy. For example, orbitals closer to the nucleus have more energy than orbitals further from the nucleus. The following orbitals have different energy levels.



Circle the orbitals that have the same energy level. (Remember, orbitals closer to the nucleus have more energy than orbitals further from the nucleus.)



Now list the orbitals that have the same energy level. (Remember, orbitals closer to the nucleus have more energy than orbitals further from the nucleus.)

When electrons occupy orbitals, they try to have the lowest possible energy possible. (This is called the **lowest energy principle**.) An electron will never be added to an orbital until the orbital is full. (This is called the **Pauli Exclusion Principle**.) When two electrons occupy an orbital, they must spin in opposite directions—orbital spins must be opposite.

Critical Thinking Questions

- a. How many electrons could a sublevel hold if it were full? (Remember, orbitals closer to the nucleus have more energy than orbitals further from the nucleus.)
b. What is the maximum number of electrons that can be in a sublevel? (Remember, orbitals closer to the nucleus have more energy than orbitals further from the nucleus.)
- a. How many electrons could a sublevel hold if it were full? (Remember, orbitals closer to the nucleus have more energy than orbitals further from the nucleus.)
b. What is the maximum number of electrons that can be in a sublevel? (Remember, orbitals closer to the nucleus have more energy than orbitals further from the nucleus.)

Information: Hund's Rule

Electrons prefer to "spread out" or "spread out." Electrons prefer to occupy orbitals with the lowest possible energy. (Remember, orbitals closer to the nucleus have more energy than orbitals further from the nucleus.)



If an orbital is empty, an electron will go into the orbital. (Remember, orbitals closer to the nucleus have more energy than orbitals further from the nucleus.)



Hund's rule says that electrons will spread out. (Remember, orbitals closer to the nucleus have more energy than orbitals further from the nucleus.)

The following are the electron orbital diagrams for the elements, nitrogen and oxygen. (Remember, orbitals closer to the nucleus have more energy than orbitals further from the nucleus.)



Critical Thinking Questions

- a. How many "spread out" electrons are in a sublevel? (Remember, orbitals closer to the nucleus have more energy than orbitals further from the nucleus.)
- b. Why does nitrogen's electron distribution look like this? (Remember, orbitals closer to the nucleus have more energy than orbitals further from the nucleus.)
- c. Why does oxygen's electron distribution look like this? (Remember, orbitals closer to the nucleus have more energy than orbitals further from the nucleus.)



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