

Experiment 10: Electrodes with a Standard Solution

From: *Basic Laboratory Experiments*, Experiment 10, 45-50

The purpose of this experiment is to determine the standard potential for several cells consisting of $\text{Zn} / \text{Zn}^{2+}$ and $\text{Cu} / \text{Cu}^{2+}$ electrodes. Electrode potentials are a half cell concept that allows us to compare standard reduction potentials relative to a common reference electrode. In this experiment, we will use a standard hydrogen electrode (SHE) as a reference electrode. Other standard electrodes, such as the silver-silver chloride electrode and the saturated calomel electrode, are often used in place of the SHE. The standard potential for a half cell is defined as the potential of the half cell relative to the SHE.

In this experiment, you will determine the standard potential for the $\text{Zn} / \text{Zn}^{2+}$ and $\text{Cu} / \text{Cu}^{2+}$ electrodes. You will also determine the standard potential for the $\text{Zn} / \text{Zn}^{2+}$ and $\text{Cu} / \text{Cu}^{2+}$ electrodes.

Objectives:

- to determine the standard potential for the $\text{Zn} / \text{Zn}^{2+}$ and $\text{Cu} / \text{Cu}^{2+}$ electrodes
- to determine the standard potential for the $\text{Zn} / \text{Zn}^{2+}$ and $\text{Cu} / \text{Cu}^{2+}$ electrodes
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- to determine the standard potential for the $\text{Zn} / \text{Zn}^{2+}$ and $\text{Cu} / \text{Cu}^{2+}$ electrodes

Materials:

- 1.0 M ZnSO_4 solution
- 1.0 M CuSO_4 solution (0.1 M)
- 1.0 M HCl
- 1.0 M NaOH
- 1.0 M KNO_3
- 1.0 M ZnSO_4 solution
- 1.0 M CuSO_4 solution
- 1.0 M HCl
- 1.0 M NaOH
- 1.0 M KNO_3

Procedure:

- 1.0 M ZnSO_4 solution
- 1.0 M CuSO_4 solution
- 1.0 M HCl
- 1.0 M NaOH
- 1.0 M KNO_3

Results:

- 1.0 M ZnSO_4 solution
- 1.0 M CuSO_4 solution

Discussion:

- 1.0 M ZnSO_4 solution
- 1.0 M CuSO_4 solution

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