

CHAPTER 7 REVIEW

Chemical Formulas and Chemical Compounds**SECTION 1****SHORT ANSWER** Answer the following questions in the space provided.

1. _____ In a Stock system name such as iron(III) sulfate, the Roman numeral tells us
 - (a) how many atoms of Fe are in one formula unit.
 - (b) how many sulfate ions can be attached to the iron atom.
 - (c) the charge on each Fe ion.
 - (d) the total positive charge of the formula unit.
2. _____ Changing a subscript in a correctly written chemical formula
 - (a) changes the number of moles represented by the formula.
 - (b) changes the charges on the other ions in the compound.
 - (c) changes the formula so that it no longer represents the compound it previously represented.
 - (d) has no effect on the formula.
3. The explosive TNT has the molecular formula $C_7H_5(NO_2)_3$.
 - _____ a. How many elements make up this compound?
 - _____ b. How many oxygen atoms are present in one molecule of $C_7H_5(NO_2)_3$?
 - _____ c. How many atoms in total are present in one molecule of $C_7H_5(NO_2)_3$?
 - _____ d. How many atoms are present in a sample of 2.0×10^{23} molecules of $C_7H_5(NO_2)_3$?
4. How many atoms are present in each of these formula units?
 - _____ a. $Ca(HCO_3)_2$
 - _____ b. $C_{12}H_{22}O_{11}$
 - _____ c. $Fe(ClO_2)_3$
 - _____ d. $Fe(ClO_3)_2$
5. _____
 - _____ a. What is the formula for the compound dinitrogen pentoxide?
 - _____ b. What is the Stock system name for the compound FeO ?
 - _____ c. What is the formula for sulfurous acid?
 - _____ d. What is the name for the acid H_3PO_4 ?

Answers Review Chemical Formulas And Chemical Compounds

Arshad Iqbal



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Table of Contents Answers Review Chemical Formulas And Chemical Compounds

1. Understanding the eBook Answers Review Chemical Formulas And Chemical Compounds
 - The Rise of Digital Reading Answers Review Chemical Formulas And Chemical Compounds
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers Review Chemical Formulas And Chemical Compounds
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Review Chemical Formulas And Chemical Compounds
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Review Chemical Formulas And Chemical Compounds
 - Personalized Recommendations
 - Answers Review Chemical Formulas And Chemical Compounds User Reviews and Ratings
 - Answers Review Chemical Formulas And Chemical Compounds and Bestseller Lists
5. Accessing Answers Review Chemical Formulas And Chemical Compounds Free and Paid eBooks
 - Answers Review Chemical Formulas And Chemical Compounds Public Domain eBooks
 - Answers Review Chemical Formulas And Chemical Compounds eBook Subscription Services
 - Answers Review Chemical Formulas And Chemical Compounds Budget-Friendly Options
6. Navigating Answers Review Chemical Formulas And Chemical Compounds eBook Formats

- ePub, PDF, MOBI, and More
 - Answers Review Chemical Formulas And Chemical Compounds Compatibility with Devices
 - Answers Review Chemical Formulas And Chemical Compounds Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers Review Chemical Formulas And Chemical Compounds
 - Highlighting and Note-Taking Answers Review Chemical Formulas And Chemical Compounds
 - Interactive Elements Answers Review Chemical Formulas And Chemical Compounds
 8. Staying Engaged with Answers Review Chemical Formulas And Chemical Compounds
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers Review Chemical Formulas And Chemical Compounds
 9. Balancing eBooks and Physical Books Answers Review Chemical Formulas And Chemical Compounds
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers Review Chemical Formulas And Chemical Compounds
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Answers Review Chemical Formulas And Chemical Compounds
 - Setting Reading Goals Answers Review Chemical Formulas And Chemical Compounds
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Answers Review Chemical Formulas And Chemical Compounds
 - Fact-Checking eBook Content of Answers Review Chemical Formulas And Chemical Compounds
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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web he concluded that at a constant pressure the volume of the gas would expand at a particular rate for each degree of temperature rise that being 1 267 per degree in 1874 victor regnault obtained better experimental results showing this number to be 1 273 and concluded that the pressure would approach zero at 1 273 degrees c

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web the volume of 1 00 mol 1 00 mol of any gas at stp standard temperature 273 15 k and pressure 1 atm is measured to be 22 414l 22 414 l we can substitute 101 325kpa 101 325 kpa for pressure 22 414 l 22 414 l for volume and 273 15 k 273 15 k for temperature into the ideal gas equation and solve for r r

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web chemists sometimes make comparisons against a standard temperature and pressure stp for reporting properties of gases 273 15 k and 1 atm 101 325 kpa 1 at stp one mole of an ideal gas has a volume of about 22 4 l this is referred to as the standard molar volume figure 9 18

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web the pressure volume temperature pvt fluids lab in ku s chemical petroleum engineering department is used to study the properties of fluids under a wide range of pressure and temperature

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web volume pressure kpa linear pressure kpa volume ml pressure kpa in part b we observed the relationship between temperature and pressure the pressure of the 125 ml flask was at 101 36 kpa when the temperature was at 23 2 c

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web the properties of water at constant volume can be represented as a function of pressure and temperature as shown in

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web figure pageindex 5 the empirically determined relationships among pressure volume temperature and amount of a gas the thermometer and pressure gauge indicate the temperature and the pressure qualitatively the level in the flask indicates the volume and the number of particles in each flask indicates relative amounts

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