Chapter 1 - Chemical Foundations

 Consider the element indium, atomic number 49, atomic mass 114.8 g. The nucleus of an atom of indium-112 contains:

- a. 49 protons, 63 neutrons, 49 electrons.
- b. 49 protons, 49 neutrons.
- c. 49 protons, 49 alpha particles.
- d. 49 protons, 63 neutrons.
- e. 49 protons, 112 neutrons.

| ANSWER: | d |
|---------|---|
| POINTS: | 1 |

2. A hypothetical element consists of two isotopes of masses 78.95 amu and 80.95 amu with abundances of 32.9% and 67.1%, respectively. What is the average atomic mass of this element?

| a. | 79.95 amu |
|----|-----------|
| b. | 78.95 amu |
| c. | 79.6 amu |
| d. | 80.3 amu |
| e. | 80.95 amu |

| ANSWER: | d |
|---------|---|
| POINTS: | 1 |

Naturally occurring element X exists in three isotopic forms: X-28 (27.977 amu, 92.23% abundance), X-29 (28.976 amu, 4.67% abundance), and X-30 (29.974 amu, 3.10% abundance). Calculate the atomic weight of X.

| a. | 28.09 amu |
|------|-----------|
| b. | 48.63 amu |
| c. | 27.16 amu |
| d. | 28,97 amu |
| 40.0 | 86.93 amu |

| ANSWER: | a a | |
|---------|-----|---|
| POINTS: | | ı |

4. The atomic mass of rhenium is 186.2. Given that 37.1% of natural rhenium is rhenium-185, what is the other stable isotope?

| a. | 183 Re |
|----|---------|
| b. | 187 Re |
| C. | 189 R.e |
| d. | 181 R.e |
| e. | 190 R.e |

| ANSWER: | 1 | 6 |
|---------|------|---|
| POINTS: | Si . | |

Naturally occurring element X exists in three isotopic forms: X-28 (27.977 amu, 92.23% abundance), X-29
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Answers For Chapter One Chemical Foundations Exercises

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