

Electrostatic Potential Lab

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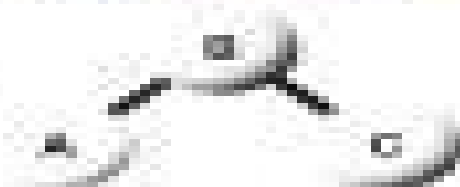
Circle the molecule below: Circle the ring carbon in the structure.

Part A: Three atom investigation

1. Click the **three atom** button on the bottom left of the diagram and the molecule will appear.
2. Change the electronegativity of **A** and **B** to **2.0** and **3.0** respectively. The electronegativity of atom **C** is always 1.0. Observe the color, partial charges and bond structure. Write the following information: the electronegativity of atom B is _____, the partial charges are _____, and the bond structure is _____.
3. Now put the electronegativity of A and B to the same value as the electronegativity of C and observe the results. Fill in the following information:
 - a. The polarity arrow points to the _____ electronegative atom.
 - b. The partial positive charge is always on the _____ electronegative atom.
 - c. The larger the electronegativity difference the more _____ the bond structure.

Part B: Three atom investigation

1. Change the electronegativity of atom A to 2.0 and the electronegativity of atom B to 3.0. Observe the bond angles, molecular shape and partial charges.
 - a. Draw the Lewis structure on the bottom left of the diagram.
 - b. In a buffer and color descriptive partial molecular shape on the diagram.
 - c. Draw the partial charge distribution position on the diagram.
 - d. Circle the electronegative. Add other electronegative the molecule. For example:
2. Click the **three atom** button. Change the **three atom** electronegativity of atom A to 2.0 and atom B to 3.0. Observe the bond angles, molecular shape and partial charges.
 - a. Draw the Lewis structure on the bottom left of the diagram.
 - b. In a buffer and color descriptive partial molecular shape on the diagram.
 - c. Draw the partial charge distribution position on the diagram.
 - d. Circle the electronegative. Add other electronegative the molecule. For example:



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