

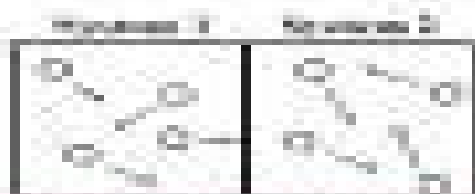
Unit 6: Thermochemistry**6.1: Heat Transfer and Thermal Equilibrium****6.2: Heat Capacity and Calorimetry****6.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann-probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with v is $\frac{1}{2}mv^2$, so the more of particles that move fast enough, the more kinetic energy we will see for a system. Therefore, the more particles move faster, the more kinetic energy we will see for the system. It's important to note that the kinetic energy is proportional to the square of the velocity.

3. Sketch a curve (graph) to represent the relative distribution of the particles in the following two gases, samples of the same gas.



- a. In System 1, is the average kinetic temperature more System 1 or System 2? Justify your answer.

System 1 is at a higher temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the particles are more compact, so they are moving faster. In System 2, the particles are more spread out, so they are moving slower. Therefore, the average KE is higher in System 1 than in System 2.

- b. In which direction will thermal energy be transferred? Justify your answer in terms of molecular collisions.

Thermal energy will be transferred from System 1 to System 2. System 1 has a higher temperature than System 2, so the particles in System 1 are moving faster. When the particles from the two systems are mixed, the faster-moving particles will collide with the slower-moving particles, transferring some of their kinetic energy to them. This will continue until the two systems reach thermal equilibrium.

Answer Key Unit 6 Thermochemistry

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