

Physics

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(Chapter 1) (Units and Measurement) (Exercises) (Class 11)

Question 1.14:

The unit of length convenient on the atomic scale is known as an angstrom and is denoted by Å: $1 \text{ Å} = 10^{-10} \text{ m}$. The size of a hydrogen atom is about 0.5 Å . What is the total atomic volume in m^3 of a mole of hydrogen atoms?

Answer 1.14:

Radius of hydrogen atom, $r = 0.5 \text{ Å} = 0.5 \times 10^{-10} \text{ m}$

Volume of hydrogen atom $= \frac{4}{3} \pi r^3$

$$= \frac{4}{3} \times \frac{22}{7} \times (0.5 \times 10^{-10})^3$$

$$= 0.524 \times 10^{-30} \text{ m}^3$$

1 mole of hydrogen contains 6.023×10^{23} hydrogen atoms.

Volume of 1 mole of hydrogen atoms

$$= 6.023 \times 10^{23} \times 0.524 \times 10^{-30}$$

$$= 3.16 \times 10^{-7} \text{ m}^3$$

Question 1.15:

One mole of an ideal gas at standard temperature and pressure occupies 22.4 L (molar volume). What is the ratio of molar volume to the atomic volume of a mole of hydrogen? (Take the size of hydrogen molecule to be about 1 Å). Why is this ratio so large?

Answer 1.15:

Radius of hydrogen atom, $r = 0.5 \text{ Å} = 0.5 \times 10^{-10} \text{ m}$

$$\text{Volume of hydrogen atom} = \frac{4}{3} \pi r^3 = \frac{4}{3} \times \frac{22}{7} \times (0.5 \times 10^{-10})^3 = 0.524 \times 10^{-30} \text{ m}^3$$

Now, 1 mole of hydrogen contains 6.023×10^{23} hydrogen atoms.

$$\therefore \text{Volume of 1 mole of hydrogen atoms, } V_a = 6.023 \times 10^{23} \times 0.524 \times 10^{-30}$$

$$= 3.16 \times 10^{-7} \text{ m}^3$$

Molar volume of 1 mole of hydrogen atoms at STP, $V_m = 22.4 \text{ L} = 22.4 \times 10^{-3} \text{ m}^3$

$$\therefore \frac{V_m}{V_a} = \frac{22.4 \times 10^{-3}}{3.16 \times 10^{-7}} = 7.08 \times 10^4$$

Hence, the molar volume is 7.08×10^4 times higher than the atomic volume. For this reason, the inter-atomic separation in hydrogen gas is much larger than the size of a hydrogen atom.

Question 1.16:

Explain this common observation clearly: If you look out of the window of a fast moving train, the nearby trees, houses etc. seem to move rapidly in a direction opposite to the train's motion, but the distant objects (hill tops, the Moon, the stars etc.) seem to be stationary. (In fact, since you are aware that you are moving, these distant objects seem to move with you).

Answer 1.16:

Line of sight is defined as an imaginary line joining an object and an observer's eye. When we observe nearby stationary objects such as trees, houses, etc. while sitting in a moving train, they appear to move rapidly in the opposite direction because the line of sight changes very rapidly.

On the other hand, distant objects such as trees, stars, etc. appear stationary because of the large distance. As a result, the line of sight does not change its direction rapidly.

Answers Physics Unit 7 Lesson 1

Jessica J Manson



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Table of Contents Answers Physics Unit 7 Lesson 1

1. Understanding the eBook Answers Physics Unit 7 Lesson 1
 - The Rise of Digital Reading Answers Physics Unit 7 Lesson 1
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers Physics Unit 7 Lesson 1
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Physics Unit 7 Lesson 1
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Physics Unit 7 Lesson 1
 - Personalized Recommendations

- Answers Physics Unit 7 Lesson 1 User Reviews and Ratings
- Answers Physics Unit 7 Lesson 1 and Bestseller Lists
- 5. Accessing Answers Physics Unit 7 Lesson 1 Free and Paid eBooks
 - Answers Physics Unit 7 Lesson 1 Public Domain eBooks
 - Answers Physics Unit 7 Lesson 1 eBook Subscription Services
 - Answers Physics Unit 7 Lesson 1 Budget-Friendly Options
- 6. Navigating Answers Physics Unit 7 Lesson 1 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers Physics Unit 7 Lesson 1 Compatibility with Devices
 - Answers Physics Unit 7 Lesson 1 Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers Physics Unit 7 Lesson 1
 - Highlighting and Note-Taking Answers Physics Unit 7 Lesson 1
 - Interactive Elements Answers Physics Unit 7 Lesson 1
- 8. Staying Engaged with Answers Physics Unit 7 Lesson 1
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers Physics Unit 7 Lesson 1
- 9. Balancing eBooks and Physical Books Answers Physics Unit 7 Lesson 1
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers Physics Unit 7 Lesson 1
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers Physics Unit 7 Lesson 1
 - Setting Reading Goals Answers Physics Unit 7 Lesson 1
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers Physics Unit 7 Lesson 1
 - Fact-Checking eBook Content of Answers Physics Unit 7 Lesson 1

- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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exist with individual atoms as their basic unit it is assumed that there is only one atom in a formula if there is no numerical
subscript on the right side of an element s symbol

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web as stated earlier an atom is the smallest constituent particle of an element which exhibits the chemical properties of an
element and also can take part in a chemical reaction atoms are extremely small and their sizes are about an angstrom big 1
1 angstrom text a circ 10 10 text m big a 10 10 m

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