

CHAPTER 7 REVIEW

Chemical Formulas and Chemical Compounds

Teacher Notes and Answers

Chapter 7 SECTION 1

SHORT ANSWER

- c
- c
- 4 elements
 - 6 oxygen atoms
 - 21 atoms
 - 4.2×10^{24} atoms
- 11 atoms
 - 45 atoms
 - 10 atoms
 - 9 atoms
- N_2O_5
 - iron(II) oxide
 - H_2SO_3
 - phosphoric acid
- True
 - True
- In general, if the anion name ends in *-ate*, the corresponding acid name will end in a suffix of *-ic*. In general, if the anion name ends in *-ite*, the corresponding acid name will end in a suffix of *-ous*.
 - In general, if the anion name ends in *-ide*, the corresponding acid name will end in a suffix of *-ic* and begin with a prefix of *hydro-*. The prefix *hydro-* is never used for anions ending in *-ate* or *-ite*.
-

Compound name	Formula
Aluminum sulfide	Al_2S_3
Cesium carbonate	Cs_2CO_3
Lead(II) chloride	PbCl_2
Ammonium phosphate	$(\text{NH}_4)_3\text{PO}_4$
Hydroiodic acid	HI

SECTION 2

SHORT ANSWER

- +4
 - +6
 - 2
 - +1
 - +6
 - +5
 - +4
 - 3
- SnCl_2
 - nitrogen(IV) oxide
- fluorine
 - 0; F_2
- tin(IV) oxide
 - SnO
- NO , NO_2
 - ClO
- N_2O_5
 - +3
 - The three oxygen atoms have oxidation states of -6 total, and because the algebraic sum of the oxidation states in a neutral compound must be zero, the two nitrogen atoms must have oxidation states of +6 total, therefore +3 each.
- nitrogen(III) oxide
- Carbon is +4 and each oxygen is -2.
 - Carbon is -4 and each hydrogen is +1.
 - Each carbon is 0, each hydrogen is +1, and each oxygen is -2.
 - Each carbon is -8/3 and each hydrogen is +1.
- Each iron is +3 and each oxygen is -2.
 - Nitrogen is +4 and each oxygen is -2.
 - Chromium is +4 and each oxygen is -2.

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