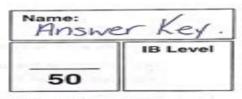
#### Chemical Quantities Unit Test



 Answer the following in the space provided. Where required, be sure to use correct scientific notation, significant digits and including proper units of measure.
 marks)

(12 marks)	is of measure.
(a) the product(s) of the reaction of $SO_{3(g)} + H_2O_{(l)} \rightarrow$	H2504(00)
(b) the product(s) of the reaction of $Ba_{(s)} + H_2O_{(t)} \longrightarrow$	Ba(OH)2000 + H219)
(c) the product(s) of the reaction of $Na_2CO_{3(s)}+HNO_{3(aq)} \rightarrow$	Nanogen + H2010 + CO2
(d) the product(s) of the reaction of $Ag_{(s)} + Pb(NO_3)_{2(aq)} \rightarrow$	no rxn.
(e) the number of moles in 9.97 x $10^{25}$ molecules of $CO_2$	166 mol
(f) the number of $H$ atoms in 2.02 moles of $CH_3(CH_2)_8COOH$	2.43 × 1025 H atoms
(g) the number of moles in a 71.92 g sample of MgSO <sub>4</sub>	0.5974 mol
(h) the mass of 3.33 x 10 <sup>20</sup> molecules of chlorine	0.03929
(i) percentage composition of ammonium phosphate chemical formula	(NH4) PO4
molar mass	149.12 A.I
payagataga composition	28.19% N 8.13% H
percentage composition	20.77% P 42.92% O

Answer the following with complete solutions, be sure to indicate all units of measure and express final answer to correct significant digits.

 A molecular compound containing the unknown element X has the formula X<sub>4</sub>O<sub>10</sub>. If it is known that the compound is 56.36% oxygen by mass, calculate the molar mass of element X and state it's identity. (2 marks)

$$56.36\% = \frac{10(16.00 \, \text{m}_{\odot})}{M_{\text{XuPo}}} \times 100\%$$
  $283.9 \, \text{m}_{\odot} = 4(M_{\text{X}}) + 10(16.00 \, \text{m}_{\odot})$   
 $M_{\text{XuPo}} = 283.9 \, \text{m}_{\odot} = 283.9 \, \text{m}_{\odot} = 30.97 \, \text{m}_{\odot} =$ 

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