

1. Answer the following in the space provided. Where required, be sure to use correct scientific notation, significant digits and including proper units of measure. (12 marks)

(a) the product(s) of the reaction of $SO_3(g) + H_2O(l) \rightarrow$

(b) the product(s) of the reaction of $Ba(s) + H_2O(l) \rightarrow$

(c) the product(s) of the reaction of $Na_2CO_3(s) + HNO_3(aq) \rightarrow$

(d) the product(s) of the reaction of $Ag(s) + Pb(NO_3)_2(aq) \rightarrow$

(e) the number of moles in 9.97×10^{25} molecules of CO_2

(f) the number of H atoms in 2.02 moles of $CH_3(CH_2)_8COOH$

(g) the number of moles in a 71.92 g sample of $MgSO_4$

(h) the mass of 3.33×10^{20} molecules of chlorine

(i) percentage composition
of ammonium phosphate

chemical formula

molar mass

percentage composition

$H_2SO_4(aq)$	
$Ba(OH)_2(aq) + H_2(g)$	
$NaNO_3(aq) + H_2O(l) + CO_2(g)$	
no rxn.	
166 mol	
2.43×10^{25} H atoms	
0.5974 mol	
0.0392 g	
$(NH_4)_3PO_4$	
149.12 $\frac{g}{mol}$	
28.19% N	8.13% H
20.77% P	42.92% O

Answer the following with complete solutions, be sure to indicate all units of measure and express final answer to correct significant digits.

2. A molecular compound containing the unknown element X has the formula X_4O_{10} . If it is known that the compound is 56.36% oxygen by mass, calculate the molar mass of element X and state its identity. (2 marks)

$$56.36\% = \frac{10(16.00 \frac{g}{mol})}{M_{X_4O_{10}}} \times 100\% \quad 283.9 \frac{g}{mol} = 4(M_X) + 10(16.00 \frac{g}{mol})$$

$$M_X = 30.97 \frac{g}{mol}$$

$$M_{X_4O_{10}} = 283.9 \frac{g}{mol}$$

$\therefore X$ is phosphorus (P)

Answer Key For Chemical Quantities Test

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Table of Contents Answer Key For Chemical Quantities Test

1. Understanding the eBook Answer Key For Chemical Quantities Test
 - The Rise of Digital Reading Answer Key For Chemical Quantities Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key For Chemical Quantities Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key For Chemical Quantities Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key For Chemical Quantities Test

- Personalized Recommendations
 - Answer Key For Chemical Quantities Test User Reviews and Ratings
 - Answer Key For Chemical Quantities Test and Bestseller Lists
5. Accessing Answer Key For Chemical Quantities Test Free and Paid eBooks
 - Answer Key For Chemical Quantities Test Public Domain eBooks
 - Answer Key For Chemical Quantities Test eBook Subscription Services
 - Answer Key For Chemical Quantities Test Budget-Friendly Options
 6. Navigating Answer Key For Chemical Quantities Test eBook Formats
 - ePub, PDF, MOBI, and More
 - Answer Key For Chemical Quantities Test Compatibility with Devices
 - Answer Key For Chemical Quantities Test Enhanced eBook Features
 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key For Chemical Quantities Test
 - Highlighting and Note-Taking Answer Key For Chemical Quantities Test
 - Interactive Elements Answer Key For Chemical Quantities Test
 8. Staying Engaged with Answer Key For Chemical Quantities Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key For Chemical Quantities Test
 9. Balancing eBooks and Physical Books Answer Key For Chemical Quantities Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key For Chemical Quantities Test
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Answer Key For Chemical Quantities Test
 - Setting Reading Goals Answer Key For Chemical Quantities Test
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Answer Key For Chemical Quantities Test

- Fact-Checking eBook Content of Answer Key For Chemical Quantities Test
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
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