

Empirical and Molecular Formula Practice

3 pages + 3-page Answer KEY + Periodic Table

EMPIRICAL AND MOLECULAR PROBLEMS

Chemistry

1 The percent composition of a certain gas is nitrogen (84.8%) and hydrogen (15.2%). Calculate the empirical formula.

Step 1: Convert % to g.

Step 3: Divide by the smallest mole value.

Step 4: Multiply by smallest number that will make a whole.

2 The compound in problem 1 has a molar mass of 44 g/mol. Calculate the molecular formula.

Step 2: Divide molecular mass (MM) by empirical mass.

Step 3: Multiply empirical formula by the answer in Step 2.

3 A compound is composed of 60.1% nitrogen and 39.9% hydrogen. What is the empirical formula of this compound?

4 The compound in problem 3 has a molar mass of 100 g/mol. Calculate the molecular formula.

5 Hydrogen, a high-energy fuel, is 14.1% hydrogen. Calculate the empirical formula.

6 Hydrogen's molar mass is 100 g/mol. Calculate the molecular formula.

7 A compound composed of hydrogen (2.14%) and oxygen (97.86%) has a molar mass of 34 g/mol. What is the empirical formula of this compound?

8 What is the molecular formula of the compound in problem 7?

9 A compound with a known molar mass of 44 g/mol contains carbon (81.7%) and hydrogen (18.3%). What is the empirical formula?

10 Calculate the empirical formula of the compound in problem 9.

ANSWER KEY
INCLUDED

Answer Key Empirical And Molecular Practice

G Thomas



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