

Acid-Base Neutralization & Titrations Practice

Name: _____

1. A 35.0 mL sample of NaOH solution is titrated to an end point by 14.8 mL 0.412M HBr solution. What is the molarity of the NaOH solution?



$$M_A V_A = M_B V_B \rightarrow M_B = \frac{M_A V_A}{V_B} \rightarrow M_B = \frac{(0.412 \text{ M})(14.8 \text{ mL})}{(35.0 \text{ mL})} \rightarrow \boxed{M_B = 0.174 \text{ M NaOH}}$$

2. If 25.0 mL of KOH were needed to neutralize 15.0 mL of 3.50M HBr, calculate the molarity of the base.



$$M_A V_A = M_B V_B \rightarrow M_B = \frac{M_A V_A}{V_B} \rightarrow M_B = \frac{(3.50 \text{ M})(15.0 \text{ mL})}{(25.0 \text{ mL})} \rightarrow \boxed{M_B = 2.10 \text{ M KOH}}$$

3. It required 34.0 mL of 1.90M KOH to neutralize 25.0 mL of H_3PO_4 . Calculate the molarity of the phosphoric acid.



$$\frac{M_A V_A}{n_A} = \frac{M_B V_B}{n_B} \rightarrow M_A = \frac{M_B V_B n_A}{V_A n_B} \rightarrow M_A = \frac{(1.90 \text{ M})(34.0 \text{ mL})(1)}{(25.0 \text{ mL})(3)} \rightarrow \boxed{M_A = 0.860 \text{ M H}_3\text{PO}_4}$$

4. A student titrated 20.0 mL of 0.500M HNO_3 with NaOH. Her data is shown in the following table.

a. Complete the following table:

	Trial 1	Trial 2	Trial 3
INITIAL VOLUME NaOH	0.00 mL	15.3 mL	30.4 mL
FINAL VOLUME NaOH	15.3 mL	30.4 mL	45.3 mL
VOLUME NaOH used	<u>15.3 mL</u>	<u>15.1 mL</u>	<u>14.9 mL</u>

b. Calculate the average volume of NaOH used from the three trials.

$$\frac{(15.3 \text{ mL}) + (15.1 \text{ mL}) + (14.9 \text{ mL})}{3} \rightarrow \frac{45.3 \text{ mL}}{3} \rightarrow \boxed{\text{Avg Volume} = 15.1 \text{ mL NaOH}}$$

c. Use your average volume of NaOH used to calculate the molarity of the base.



$$M_A V_A = M_B V_B \rightarrow M_B = \frac{M_A V_A}{V_B} \rightarrow M_B = \frac{(0.500 \text{ M})(20.0 \text{ mL})}{(15.1 \text{ mL})} \rightarrow \boxed{M_B = 0.662 \text{ M NaOH}}$$

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