

# Atomic Theory, Structure & Electron Configuration: Skill Review

**Skill: Theory & Structure** - Student can use vocab related to atomic structure for isotope table completion

17. Complete the following table:

Name	Symbol	Atomic #	Protons	Neutrons	Electrons	Mass #
Iron	$^{56}_{26}\text{Fe}$	26	26	30	26	56
Selenium		34	34	48	34	78
Barium	$^{137}_{56}\text{Ba}$	56	56	81	56	137
Uranium		92	92	143	92	235

**Skill: Electron Configuration** - Student can define vocabulary related to electron configurations

18. Provide a definition or explanation for writing an electron configuration:

An electron configuration is similar to writing down an "address" for the highest likelihood of locating of an electrons. The electron configurations use a code with numbers representing principle energy levels, the letters s, p, d & f to represent sublevels and subscript values to represent the number of electrons in each sublevel.

**Skill: Theory & Structure** - Student understands Dalton's contribution to atomic theory

19. What were the 3 parts to Dalton's atomic theory?

1. All elements are made of atoms
2. All atoms of an element are identical
3. All atoms of different elements are different
4. Law of constant composition
5. Law of conservation of matter

**Skill: Electron Configuration** - Student can write electron configurations & orbital diagrams

20. Write the FULL electron configuration for each of the following elements:

- a. Ne -  $1s^2 2s^2 2p^6$
- b. V -  $1s^2 2s^2 2p^6 3s^2 3p^4 4s^1 3d^1$
- c. Ag -  $1s^2 2s^2 2p^6 3s^2 3p^6 4s^2 3d^{10} 4p^6 5s^1 4d^9$

21. Write the orbital (box) diagram and electron configuration for:



**1s<sup>2</sup> 2s<sup>2</sup> 2p<sup>6</sup> 3s<sup>2</sup> 3p<sup>4</sup> 4s<sup>1</sup> 3d<sup>5</sup>**  
 \*If the class is unaware of exceptions to electron configurations & orbital diagrams in transition metals, an acceptable orbital diagram prediction could end in  $4s^2 3d^4$

# Answer Key Atomic Theory Unit Regents Review

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