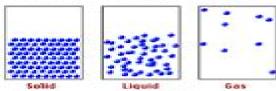
Kinetic Molecular Theory Worksheet

 The word kinetic comes from a Greek word that means "to move." The kinetic molecular theory is based upon the assumption that particles of matter (atoms or molecules) are in constant random.

Of the three states of matter, which one has the most kinetic energy? Gas.

motion.

Which state of matter has particles that are separated by the largest distance? <u>Gas</u>



- 4. A scientific theory is an explanation of some type of natural phenomena. Theories are normally developed from careful study of the way the world behaves. Let's look at how gases behave and see if the kinetic molecular theory makes sense.
- Compared to liquids and solids, gases tend to have <u>lower</u> densities. This can be explained because the particles of gas are <u>in constant motion</u>.
- 6. If you apply pressure to a sample of gas, it is fairly easy to compress its volume (think about what would happen to a balloon if you squeeze it gently). This can be explained because there is a lot of empty space between gas particles.
- 7. If someone sprays perfume in one corner of the room, eventually a person on the other side of the room can smell it. This can be explained because gas particles move <u>quick</u> and <u>random</u>. In general, we would expect lighter gas particles to travel <u>faster</u> than heavier gas particles.
- 8. When two gases mix together or move through each other, this process is known as diffusion. When gas particles escape out of a tiny hole in a container, this process is known as effusion. You should know the difference between these two words so you can avoid any confusion!
- 9. Kinetic molecular theory can be summarized as follows:
 - Gas particles are in constant motion.
 - b. Gas particles are separated by relatively large distances.
 - When gas particles collide, they do not transfer kinetic energy.
 - d. Gas particles have no attractive or repulsive forces between them.
 - The kinetic energy of a gas is dependent on the temperature of the gas.

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