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Calculating pH

How is pH related to the concentration of hydronium ions?

In biology and other science courses pH is introduced as a way to quantify the acidity or basicity of a solution. This property can be measured using a pH probe or with an indicator paper strip that changes color at different pH values. But, what is actually being measured? We know that a pH of 7 is neutral, below 7 is acid, and above 7 is base, but why? What in the solution is the paper strip or probe actually reacting with?

Model 1 - Ion Concentrations for Acids and Bases

| Beaker | Solution | Acidic, Basic or Neutral? | [H ₃ O*] | [OH-] | [H ₃ O*] × [OH: |
|--------|-----------------------|------------------------------|-------------------------|-------------------------|----------------------------|
| 1 | 0.10 M HCl(aq) | Acidic | 1.0×10 ⁻¹ M | 1.0×10 ⁻¹³ M | 1×10-11 |
| 2 | 0.0010 M HCl(aq) | Acidic | 1.0×10 ⁻⁵ M. | 1.0×10 ⁻¹¹ M | I×10-14 |
| 3 | 0.000010 M HCl(aq) | Acidic | 1.0×10 ⁻⁵ M | 1.0×10⁻º M | $I \times IO^{-td}$ |
| 4 | 0.0000010 M HCl(aq) | Acidic | 1.0×10 ⁻⁶ M | 1.0×10 ⁻⁴ M | $I \times IO^{-t4}$ |
| 5 | 0.00000010 M HCl(aq) | Neutral | 1.0×10 ⁻³ M | 1.0×10 ⁻³ M | $I \times IO^{-14}$ |
| - 6 | 0.00000010 M NaOH(aq) | Neutral | 1.0×10 ⁻² M | 1.0×10 ⁻⁷ M | 1×10-14 |
| 7 | 0.0000010 M NaOH(aq) | Basic | 1.0×10 ⁻⁸ M | 1.0×10 ⁻⁴ M | 1×10-34 |
| 8 | 0.00010 M NaOH(aq) | Basic | 1.0×10 ⁻¹⁰ M | 1.0×10 ⁻⁴ M | I×10-14 |
| 9 | 0.010 M NaOH(aq) | Basic | 1.0×10 ⁻¹² M | 1.0×10 ⁻² M | 1×10-14 |
| 10 | 0.10 M NaOH(aq) | Basic | 1.0×10 ⁻¹³ M | 1.0×10 ⁻¹ M | 1×10-14 |

1. What does the symbol [H₃O*] in Model 1 indicate?

The molar concentration of hydronium ion.

- In Beaker 2, which ion has a higher concentration, hydronium ion or hydroxide ion?
 Hydronium ion.
- Describe how the concentration of hydronium ion was calculated for Beaker 3 in Model 1 from the concentration of the acid.

Since HCl is a strong acid, all of the molecules ionize in water. Therefore, the H₃O' concentration is equal to the initial acid concentration in the solution.

 Describe how the concentration of hydroxide ion was calculated for Beaker 8 in Model 1 from the concentration of the base.

Since NaOH is a strong base, it completely disorciates in toater. Therefore, the OH concentration is equal to the initial base concentration in the solution.

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Joacim Rocklöv

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