

Hess law lab

The objective of the present experiment is to determine ΔH_1 at constant T and P for the reaction:



To do this, the ΔH 's of two reactions will be measured along with a ΔH value taken from the literature. Specifically the reactions are:



If the temperature is 25.0 °C and the pressure is 1 atm,

$$\Delta H_4 = 2 \Delta H_f^\circ (\text{H}_2\text{O(l)}) = 2 (-285.6 \text{ kJ/mol}) = -571.6 \text{ kJ/mole}$$

where the enthalpy of formation of liquid water is defined as



Equations (2), (3), and (4) can be manipulated to obtain equation (1). Specifically,

$$\Delta H_1 = 2 \Delta H_2 - 2 \Delta H_3 + \Delta H_4 \quad (5)$$

The ΔH_1 obtained from (5) will be compared to the accepted value and a percent difference will be reported.

The experiment was performed as follows. For the Mg + HCl reaction, 50.0 mLs of 1M HCl solution was added to two nested Styrofoam cups. A thermometer was placed in a hole in a plastic lid which covered the Styrofoam cup, to minimize heat loss to the air. The starting temperature of the HCl solution was recorded. Approx. 0.12g of Mg was added to the HCl solution, the thermometer/lid was quickly replaced, and the temperature was monitored. The maximum temperature was recorded as the final temperature. The procedure for the MgO + HCl reaction was identical, except that approx. 0.50g of MgO was used instead of Mg.

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Line 1: 0.113

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