

Chemical Equations and Reactions

QUIZ
CH #8
SECTION #1

Section Quiz: Describing Chemical Reactions

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

B

1. Which of the following is true regarding the law of conservation of mass?
- It dictates that the number of molecules on each side of a chemical equation must be the same.
 - It dictates that the number of atoms of each element must be the same on both sides of a chemical equation.
 - It states that the mass of the reactants must remain constant in order for a chemical reaction to proceed.
 - It does not apply to chemical reactions.

D

2. Which of the following may indicate that a chemical reaction has occurred?
- release of energy as light
 - a color change
 - gas bubble formation
 - All of the above

D

3. Which of the following is a product in the reaction described by the word equation below?



- iron(II) sulfate
- copper
- iron
- Both (a) and (b)

C

4. What does CH_4 represent in the balanced chemical equation shown below?



- one atom of CH_4
- five atoms of CH_4
- one molecule of CH_4
- 1 g of CH_4

↑
1C

5. Which of the following is not true regarding a properly written chemical equation?
- It must be balanced.
 - It must contain proper chemical formulas.
 - It must show the formation of a precipitate or a gas.
 - It should represent known data.

Answer Key To Section 1describing Chemical Reactions

D Keegan



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