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web  $m \times c \times \Delta t$  equation 1 where m is the total mass of the solution solute plus solvent c is the specific heat of the solution and  $\Delta t$  is the observed temperature change the specific heat of the solution is generally assumed to be the same as that of *guided inquiry designing the ideal hand warmer* - Dec 27 2022

web the heat of reaction  $\Delta H_{\text{soln}}$  is written after the products in units of kJ mol rxn for example sodium hydroxide dissolves exothermically releasing 44.2 kilojoules per mole dissolved this process is represented as  $\text{NaOH (s)} \rightarrow \text{NaOH (aq)}$   $\Delta H_{\text{soln}} = -44.2 \text{ kJ mol rxn}$  write equations to similarly represent the dissolving process for

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web when the hand warmer pack is activated the solid dissolves in water and produces a large temperature change in the form of heat in this laboratory experiment the heat of solution will be calculated from the data for six salts in order to

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web background the hand warmers are filled with sodium acetate click the metal disc and you will notice the liquid changes state and becomes a solid in the process heat is released the sodium acetate is a supersaturated solution

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web oct 9 2021 sodium acetate coin or other small piece of metal plastic bags seal the coin inside a small bag place this bag inside a larger bag containing supersaturated sodium acetate get this by dissolving as much sodium acetate as possible in hot water this chemical hand warmer is reusable

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web the goal of my experiment and research is to evaluate the various number of compounds to be used as hand warmers and locate the ones that can raise the temperature to 20 degrees celsius in 100 g water and costs as little as possible

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web in this lab you are challenged to use chemistry to determine which ionic compound makes the best hand warmer in this experiment you will collect data that will allow you to calculate the heat of solution  $\Delta H_{\text{soln}}$  or the amount of energy released or absorbed in kJ/mol of the substance being dissolved

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web mar 4 2014 contains water the hand warmer is struck in a manner that ruptures the inner pouch releasing the ionic salt into the water of the outer pouch the salt dissolves and the water warms this type of hand warmer tends to produce a more vigorous heat than the dry powder type of hand warmer but does not produce heat for quite as long

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web hand warmers provide a unique and fun way to study the chemistry of crystal formation and heat generation by pressing a button in a pouch which contains a supercooled solution you start a rapid exothermic heat producing crystallization in this science fair project you will determine how the starting temperature affects hand warmer chemistry

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web lab investigation in today s experiment you will design and execute an experimental procedure to determine which of three ionic compounds is most suitable for use in a hand warmer the three compounds you will use is determined by the group number you are assigned the ideal hand warmer

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