

SCH4U Unit 4 - Quiz – Equilibrium

Name: _____ Period: _____

Q1: (10)

Q2: (10)

Q3: (10)

Q4: (10)

Q5: (10)

Write the formulas for:

$$K_{eq} = \frac{[products]}{[reactants]} \quad ; \quad K_{eq} = \frac{P_{products}}{P_{reactants}}$$

$$K_{eq} = \frac{P_{CO_2}^2 P_{H_2O}}{P_{CH_4} P_{H_2O}^2} \quad ; \quad K_{eq} = \frac{P_{CO_2}^2 P_{H_2O}}{(P_{CH_4}) (P_{H_2O})^2}$$

Q6: (10) Q7: (10) Q8: (10) Q9: (10) Q10: (10)

Q11: (10) Q12: (10)

Question 13:

13. Hydrogen sulfide, $H_2S(g)$, is a slightly poisonous gas. At 1000°C, an equilibrium mixture contains 0.020 moles of hydrogen, $H_2(g)$, 0.010 moles of hydrogen sulfide, with an equilibrium pressure of 0.600 atm. The reaction is:



If the value of K_{eq} is 0.00010 at 1000°C, what concentration of $H_2(g)$ is present at equilibrium?

a. 0.00010

b. 2.0×10^{-4} moles/L

c. 1.0×10^{-4} moles/L

d. 1.0×10^{-2} mol

e. Determine the value of K_{eq} at equilibrium.

f. 0.00010

What other species have chemical formula consisting of 2 hydrogen atoms alone?

g. Hydrogen

	$2H_2(g)$	$S_2(g)$	$2H_2S(g)$
Initial (moles present)			
Change (moles present)			
Final (moles present)	0.020	0.010	0.0

f. Set up your equations:

Answers For Sch4u Unit 4

Lei Shi



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