

QUESTION 2

(a) (i) The first ionization energy of an element is the amount of energy required to remove one electron from each atom in a mole of gaseous atoms, producing one mole of gaseous ions with a positive charge.

(ii) The electronic Configuration of "Be" is $1s^2 2s^2$ and that of "B" is $1s^2 2s^2 2p^1$. All the subshells of Beryllium is filled and there is some special or extra energetic stability associated with that. This means that Beryllium is more stable than boron and hence has a higher ionization energy.

(b)

$$\frac{e \times F}{It} = \frac{M}{m}$$

$$\frac{e \times 96500}{0.12 \times 500} = \frac{48}{0.015}$$

Cross-multiplying

$$e = \frac{0.12 \times 500 \times 48}{96500 \times 0.015}$$

$$e = 1.99$$

$$e \approx 2$$

The charge is +2

(c) (i) Pure alumina in molten cryolite, Na_2AlF_6

- (ii) • Availability of raw materials
- Availability of electric power

(d) (i) Paramagnetism is a property exhibited by ferrous atoms or ions having one or more unpaired electrons which are attracted into a magnetic field. e.g. Iron atom has four unpaired 3d electrons.

(d)(ii) Fe^{2+}
I) $2d^6$
2d

II) $2d^5$

III) $2d^4$
6
2d

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Hongru Du



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1. Understanding the eBook Answers Of Chemistry 2 Essay And 1 Obj
 - The Rise of Digital Reading Answers Of Chemistry 2 Essay And 1 Obj
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers Of Chemistry 2 Essay And 1 Obj
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Of Chemistry 2 Essay And 1 Obj
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers Of Chemistry 2 Essay And 1 Obj
 - Personalized Recommendations
 - Answers Of Chemistry 2 Essay And 1 Obj User Reviews and Ratings
 - Answers Of Chemistry 2 Essay And 1 Obj and Bestseller Lists
5. Accessing Answers Of Chemistry 2 Essay And 1 Obj Free and Paid eBooks
 - Answers Of Chemistry 2 Essay And 1 Obj Public Domain eBooks
 - Answers Of Chemistry 2 Essay And 1 Obj eBook Subscription Services

- Answers Of Chemistry 2 Essay And 1 Obj Budget-Friendly Options
- 6. Navigating Answers Of Chemistry 2 Essay And 1 Obj eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers Of Chemistry 2 Essay And 1 Obj Compatibility with Devices
 - Answers Of Chemistry 2 Essay And 1 Obj Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers Of Chemistry 2 Essay And 1 Obj
 - Highlighting and Note-Taking Answers Of Chemistry 2 Essay And 1 Obj
 - Interactive Elements Answers Of Chemistry 2 Essay And 1 Obj
- 8. Staying Engaged with Answers Of Chemistry 2 Essay And 1 Obj
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers Of Chemistry 2 Essay And 1 Obj
- 9. Balancing eBooks and Physical Books Answers Of Chemistry 2 Essay And 1 Obj
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers Of Chemistry 2 Essay And 1 Obj
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers Of Chemistry 2 Essay And 1 Obj
 - Setting Reading Goals Answers Of Chemistry 2 Essay And 1 Obj
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers Of Chemistry 2 Essay And 1 Obj
 - Fact-Checking eBook Content of Answers Of Chemistry 2 Essay And 1 Obj
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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