

CHAPTER 10: Reactions

Reaction Energy

SECTION 1

Write a balanced chemical equation for the following reactions in the space provided.

- For elements in their standard state, the value of ΔH_f° is _____.
- The formation and decomposition of a substance is represented by the following thermodynamic equations:



- Is the energy being added to or is it being subtracted from the decomposition?
- What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

- _____ $\Delta H_{\text{f}}^\circ$ of a reaction is 2074 kJ whereas $\Delta H_{\text{d}}^\circ$ of a reaction is 1037 kJ . What is the $\Delta H_{\text{f}}^\circ$ of the reaction?

- _____ A chemical bond is a specific bond of H_2O (Fig. 10.1). The $\Delta H_{\text{f}}^\circ$ of $\text{H}_2\text{O}(l)$ is -285.8 kJ/mol . The $\Delta H_{\text{d}}^\circ$ of $\text{H}_2\text{O}(l)$ is 285.8 kJ/mol . What is the $\Delta H_{\text{f}}^\circ$ of the reaction?

- The products is a chemical bond of H_2O (Fig. 10.1). The $\Delta H_{\text{f}}^\circ$ of $\text{H}_2\text{O}(l)$ is -285.8 kJ/mol .

- What is the value of $\Delta H_{\text{f}}^\circ$ for this reaction?

Answers For Reaction Energy Section 1

Ying Liu



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