Extra Practice

Chapter 4

Lessons 4-1 and 4-2

Gragoli maetic revolucion.

Low on Section





3. p = 2x2



4. p = (a - - 1)





 $-200x - 25^{2} + 3$



identify the vertex, axis of symmetry, minimum or maximum value, and domain and range of each function.

 $T_{\rm c}/y = 40x - 20^{\circ}$ 62, 60; x - 2; minimum 0; domain: all real numbers: managers on in 18

 $4c \cdot f(x) = (x + 1)^2 + 2$ (-4, 2)(x) = -4; entirelessors 2: dominion all real numbers: Charles Hall in Z.

 $\Phi_{i} : p_{i} = -10^{\circ} p_{i} = 40^{\circ} = 10^{\circ}$ 64: - 25: a - 4: manifesteriore - 2: domain: all real roundwirt; runger; print -2.

10. Alab - at - ac + 5. GP., 10: at - 2: minimum; to domain; all real numbers: managers of Co. R.

11. f(x) = -2x2 = 4x - 3 disensing all real mainteens.

 $32. y = x^2 + 5x - 34$ $\{X_i - X\}_i = 1$; maximum: $\{X_i - X_i\}_{i=1}^{n} = 1$; minimum: 1 domein all real pumbers. managem to be with

the ground a of the ball after I seconds can be modeled by the equation print with the Print and the second

manager year. - 5

a. What is the y-intercept of the equation? 100

The y-intercept is the starting. Describe the meaning of the p-intercept of the graph of the equation, beight of the ball.

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KJ Lindholm-Leary

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apr 25 2017 this is understood when you consider the two chemicals appear unchanged before and after the reaction therefore the net ionic equation is 2cl aq 2ag aq

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ions that remain essentially unchanged during a reaction are called spectator ions this means you can ignore them when you write the ionic equation you only need to model how the solid

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the first step to writing a net ionic equation is balancing the chemical equation present let's use the reaction between sodium chloride and silver nitrate as an example next we write the

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highlights learning objectives by the end of this section you will be able to derive chemical equations from narrative descriptions of chemical reactions write and balance chemical

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this net ionic equation indicates that solid silver chloride may be produced from dissolved chloride and silver i ions regardless of the source of these ions these molecular and

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write the overall chemical equation the complete ionic equation and the net ionic equation for the reaction of aqueous silver fluoride with aqueous sodium phosphate to give solid silver

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the complete ionic equation for this reaction is as follows equation 4 13 2ag aq 2f aq 2nh4 aq cr2o72 aq ag2cr2o7 s 2nh4 aq 2f aq because two nh 4 aq

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