



Suggested answers to in-text activities and unit-end exercises

Topic 6 Unit 24

In-text activities

Checkpoint (page 38)

- 1 a) P–H non-polar
 $\delta^- \quad \delta^+$
 b) H–F
 $\delta^- \quad \delta^+$
 c) O–N
 $\delta^- \quad \delta^+$
 d) F–F non-polar
 $\delta^- \quad \delta^+$
 e) N–H
 $\delta^- \quad \delta^+$
 f) C–I non-polar

- 2 In NH_3 , the bond pairs of electrons are attracted towards the nitrogen atom to a greater extent as nitrogen is more electronegative than hydrogen.

The bond pairs of electrons repel each other to a greater extent and thus the H–N–H bond angle is greater.

In NF_3 , the bond pairs of electrons are closer to the fluorine atom as fluorine is more electronegative than nitrogen.

The bond pairs of electrons repel each other to a less extent and thus the F–N–F bond angle is smaller.

Checkpoint (page 44)

Molecular formula	Electron diagram	Shape of molecule	Polar bond	Polar molecule?
H_2O	$\begin{array}{c} \text{H} \times \times \text{O} \times \times \\ \times \times \quad \times \times \\ \text{H} \end{array}$	V-shaped	O–H	yes
BeH_2	$\text{H} \times \times \text{Be} \times \times \text{H}$	linear	Be–H	no
CH_2F_2	$\begin{array}{c} \text{H} \\ \times \times \quad \times \times \\ \text{H} \times \times \text{C} \times \times \text{F} \times \times \\ \times \times \quad \times \times \\ \times \times \quad \times \times \\ \text{F} \end{array}$	tetrahedral	C–F	yes
Cl_2O	$\begin{array}{c} \times \times \quad \times \times \\ \times \times \text{Cl} \times \times \text{O} \times \times \\ \times \times \quad \times \times \\ \times \times \quad \times \times \\ \text{Cl} \end{array}$	V-shaped	Cl–O	yes

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Terry C. Jones



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