

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula from a chemical structure without performing stoichiometric calculations.
2. To calculate the number of moles of a solid from a weighing measurement in the lab.
3. To determine the molar mass of a solid.
4. To determine the theoretical amount of product that would be produced.
5. To compare the theoretical amount to the actual amount of product formed and calculate the percent yield.

Reaction equation:

Reaction of sodium chloride solution: 25.0 mL

Reaction of calcium chloride solution: 25.0 mL

Reaction of 0.1M NaOH: 5.00 g

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Procedure:

Calculations and Results:

1. Calculations:

$$\text{moles of sodium chloride} = 0.1 \text{ M} \times 0.025 \text{ L} = \frac{0.1}{1000} \times 25.0 = 0.0025 \text{ mol}$$

$$\text{moles of calcium chloride} = 0.1 \text{ M} \times 0.025 \text{ L} = \frac{0.1}{1000} \times 25.0 = 0.0025 \text{ mol}$$

$$\text{moles of sodium hydroxide produced} = 0.0025 \text{ mol} \times 2 = 0.005 \text{ mol}$$

$$\text{mass of product as formed} = 0.005 \text{ mol} \times 40.0 \text{ g/mol} = 0.200 \text{ g}$$

2. Chemical equation: $\text{Na}_2\text{CO}_3 + \text{CaCl}_2 \rightarrow \text{CaCO}_3 + 2\text{NaCl}$

3. Reaction equation:

Reaction of 0.1M NaOH: 5.00 g

Reaction of 0.1M NaOH: 5.00 g

$$\text{moles of NaOH} = \frac{5.00 \text{ g}}{40.0 \text{ g/mol}} = 0.125 \text{ mol}$$

Answers For Lab 6b Heath Chemistry

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Answers For Lab 6b Heath Chemistry:

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