

Examples of Covalent Bonding



4 unpaired
e⁻, can
form 4
bonds



1 unpaired
e⁻, can form
1 bond



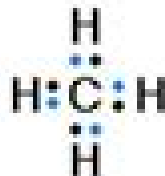
3 unpaired e⁻,
can form 3
bonds



2 unpaired e⁻,
can form 2
bonds



1 unpaired
e⁻, can form
1 bond



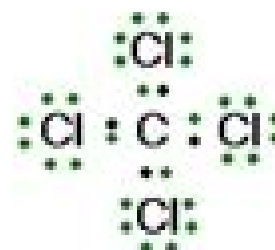
methane



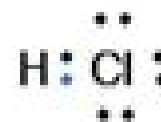
diatomic,
molecular
chlorine



hydrogen
sulfide



tetrachloro
methane



hydrogen
chloride

Unpaired electrons from each atom pair up with the unpaired electrons from the other atom. The shared pairs of electrons are covalent bonds. If the bonds are between atoms with the same electronegativity (like in the Cl-Cl bond), the bond is **NONPOLAR** covalent. If the bonds are between atoms with different electronegativities (all of the rest of the bonds), then the bond is **POLAR** covalent.

Mark Rosengarten Bonding Homework

L Reisser



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