- 16. Increasing the temperature generally____
 - decreases the solubility of a solid in a liquid but increases the solubility of a gas in a liquid.
 - B. increases the solubility of a solid in a liquid but decreases the solubility of a gas in a liquid.
 - increases the solubility of both a solid and a gas in a liquid.
 - decreases the solubility of both a solid and a gas in a liquid.
- A white precipitate was formed when BaCl_{2(ag)} was added to an aqueous solution of a salt X. The precipitate dissolved in dilute HCl with rapid efferyescence. Salt X is likely to contain
 - A. SO₄2- ions.
 - B. SO₃²⁻ ions.
 - C. NO₃^{*} ions.
 - D. CO₃²⁻ ions.
- Before a reaction could take place, there should be
 - A. ionization of reactant particles.
 - B. breakage of bonds of products.
 - C. breakage of bonds of reactants.
 - D. ionization of product particles.
- 19. Consider the following reaction equation:

Silicon(IV) oxide is acting as

- A. a basic oxide.
- B. a reducing agent.
- c. an acidic oxide.
- D. an oxidizing agent.

- 20. Which of the following acids would form normal salt only?
 - A. Tetraoxosulphate(VI) acid
 - B. Trioxonitrate(V) acid
 - C. Tetraoxophosphate(V) acid
 - D. Trioxosulphate(IV) acid
 - What is the partial pressure of oxygen at s.t.p. in a gaseous mixture containing 100 cm³ of oxygen and 900 cm³ of nitrogen gas?
 - A. 1.0 atm
 - Ek. 0.1 anton
 - C. 9.0 atm
 - D. 0.9 atm
 - Graphite is used as a dry lubricant due to the presence of
 - A. mobile electrons.
 - B. free electrons.
 - C. octahedral structures.
 - D. layered structures.
- 23. Which of the following gases is alkaline?
 - A. NO
 - B. CO₂
 - C. NH₃
 - D. N₂O
- 24. Which of the following statements about an equilibrium system is correct?
 - Forward and backward reactions occur at the same rate.
 - The concentrations of reactants must equal that of the products.
 - The concentrations of reactants and products can be changed by adding a catalyst.
 - D. The concentrations of reactants and products are not affected by a change in temperature.

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