

Chemistry, Part 1

Module 14 Notes



Big Question

How are metals used in chemical reactions? and produced by chemical reactions?

Big Idea

Using chemical analysis, you can identify the metals of a substance or pure metal and other substances. A metal is often used in manufacturing. It is used in construction or industrial analysis with chemical reactions.

Learning Objectives

After completing this module, students will be able to:

- Use the law of conservation of mass to balance chemical reactions.
- Use chemical reactions to describe processes of materials.
- Compare different metals and reactions.
- Describe the uses of different metals.

Module 14 Objectives

Learning Objective Write and balance chemical equations.

Fill in the blanks.

1. $2\text{H}_2 + \text{O}_2 \rightarrow 2\text{H}_2\text{O}$ (Balanced)
2. $\text{Fe} + \text{CuSO}_4 \rightarrow \text{FeSO}_4 + \text{Cu}$ (Balanced)
3. $\text{Ca} + \text{Mg} \rightarrow \text{Ca} + \text{Mg}$ (Not balanced)

Chemical Reactions and Equations

Write the chemical equation for the reaction of iron and oxygen.

Chemical equation: $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$ (Balanced)

Chemical equation: $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$ (Balanced)

Chemical equation: $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$ (Balanced)

Chemical equation: $\text{Fe} + \text{O}_2 \rightarrow \text{Fe}_2\text{O}_3$ (Balanced)

Module 14

Learning Objective Write and balance chemical equations.

Fill in the blanks.

1. 1 mole of iron reacts with 1 mole of oxygen to form 1 mole of iron(III) oxide.
2. The mass of iron(III) oxide is 160 g/mol. The mass of iron is 56 g/mol. The mass of oxygen is 16 g/mol.

Answers For Chemistry Module 14

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