

CHAPTER 10: Reactions

Reaction Energy

SECTION 1

Write a balanced chemical equation for the following reactions in the space provided.

- For elements in their standard state, the value of ΔH_f° is _____.
- The formation and decomposition of a substance is represented by the following thermodynamic equations:



- Is the energy being added to or is it being subtracted from the decomposition?
- What is the correct sign for the enthalpy change in this decomposition reaction?

Write the correct answer in the space to the left. Show all your work in the space provided.

- _____ $\Delta H_{\text{f}}^\circ$ of a substance is 207 kJ/mol . How much energy is released when 1 mole of this substance is formed?

- _____ A substance has a specific heat of $0.001 \text{ J/g}^\circ\text{C}$. If 1 mole of this substance is heated from 0°C to 100°C , how much energy is required to raise the temperature of 1 mole of this substance from 0°C to 100°C ?

- The product is a molecular compound and the reactant is a mixture, and the reaction is exothermic. Is this correct?

- What is the value of ΔH for this reaction?

Answers For Reaction Energy Section 1

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