

Chemical Reactions

Section 9.1 Reactions and Equations

In your textbook, read about evidence of chemical reactions.

For each statement, write *yes* if evidence of a chemical reaction is present. Write *no* if there is no evidence of a chemical reaction.

- | | |
|---------------|---|
| <u>yes</u> | 1. A tomato smells rotten. |
| <u>no</u> | 2. A drinking glass breaks into smaller pieces. |
| <u>no</u> | 3. A piece of ice melts. |
| <u>yes</u> | 4. Drain cleaner is mixed with water and the solution becomes warm. |
| <u>yes</u> | 5. Candle wax burns. |
| <u>yes/no</u> | 6. Molten candle wax solidifies. |
| <u>yes</u> | 7. Green leaves turn yellow and red as the seasons change. |
| <u>yes</u> | 8. Baking powder produces a gas that makes a cake rise. |

In your textbook, read about how to represent chemical reactions and how to balance chemical equations.

Use the terms below to complete the passage. Each term may be used once, more than once, or not at all.

arrow	plus sign	(s)	(l)
reactant	product	(g)	(aq)

The fuel for the space shuttle is hydrogen, which burns in oxygen to produce water vapor and energy. In this chemical reaction, hydrogen is a(n) **(9)** reactant, oxygen is a(n) **(10)** reactant, and water vapor is a(n) **(11)** product. In a chemical equation for this reaction, a(n) **(12)** arrow is used to separate hydrogen and oxygen from water vapor and energy. A(n) **(13)** plus sign is used to separate the symbols for hydrogen and oxygen. A(n) **(14)** (g) symbol is used to tell the state of hydrogen in the reaction, a(n) **(15)** (g) symbol is used for the state of oxygen, and a(n) **(16)** (g) symbol is used for the state of water vapor.

Answer Key To Study Packet Chemical Reactions

James E. Brady, Fred Senese



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A Life Scientist's Guide to Physical Chemistry Marc R. Roussel, 2012-04-05 Demonstrates how the tools of physical chemistry can be applied to biological questions with numerous exercises and clearly worked examples

Study Guide to Accompany Roach's Introductory Clinical Pharmacology Susan M. Ford, Sally S. Roach, 2009-10-01 This Study Guide is designed to accompany the Ninth Edition of Roach's Introductory Clinical Pharmacology

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Spotlight Science Teacher Support Pack 9 Keith Johnson, Lawrie Ryan, Sue Adamson, 2004 This Framework Edition Teacher Support Pack offers support and guidance

Chemical Reactions, Materials and Particles Gill Murphy, 2006-03 Specials are a teaching resource designed for KS3 students whose literacy skills are considerably lower than their age These books have an older format to counteract the simple text and cover various topics They include activities visuals and assessment sheets as well as teacher pages

Fundamentals of Ordinary Differential Equations Mohit Chatterjee, 2025-02-20 Fundamentals of Ordinary Differential Equations is a comprehensive guide designed for students researchers and professionals to master ODE theory and applications We cover essential principles advanced techniques and practical applications providing a well rounded resource for understanding differential equations and their real world impact The book offers a multifaceted approach from basic principles to advanced concepts catering to fields like physics engineering biology and economics Mathematical ideas are broken down with step by step explanations examples and illustrations making complex concepts accessible Real world examples throughout each chapter show how ODEs model and analyze systems in diverse disciplines We also explain numerical methods such as Euler s method Runge Kutta and finite differences equipping readers with computational tools for solving ODEs Advanced topics include bifurcation chaos theory Hamiltonian systems and singular perturbations providing an in depth grasp of ODE topics With chapter summaries exercises glossaries and additional resources Fundamentals of

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