

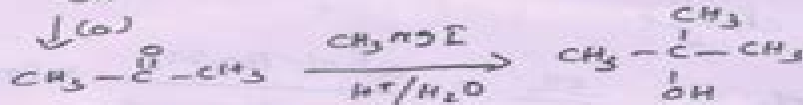
(51)

Pb^{+2} is more stable than Pb^{+4}
 F_2Cl changes to Pb^{+2} to Pb^{+4}

(52)



(53)



(54)

1 Sr^{+2} creates one cationic vacancy

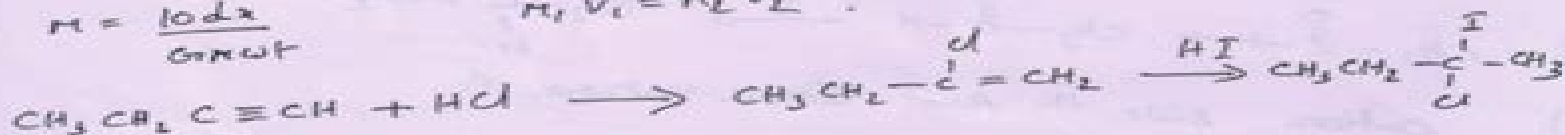
$$\text{no of cationic vacancies} = \frac{10^{-4}}{100} \times 6 \times 10^{23} = 6 \times 10^{17}$$

(55)

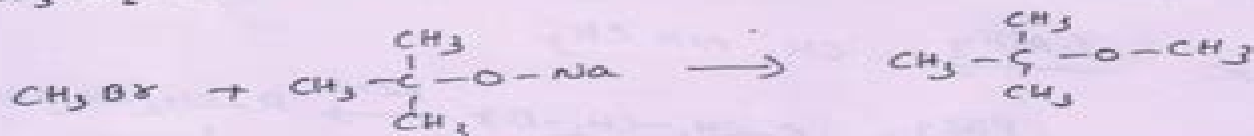
$$M = \frac{10dx}{G \cdot m \cdot W}$$

$$M_1 V_1 = M_2 V_2$$

(56)



(57)



(58)

The non metal which R.P can displace low R.P non metal from their salt solution.

(60)



(61)

IIA group carbonates more stable than IIA group carbonates.

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