

CHEMISTRY MULTIPLE CHOICE QUESTIONS

A. Atoms, molecules and stoichiometry

2002 -2019



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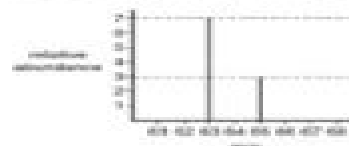
As a simplification, an adult human can be considered to have a daily diet of 1.80 kg of carbohydrate (empirical formula $C_6H_{12}O_6$).

Which mass of carbon dioxide does a person produce each day if all the carbohydrate eaten is digested and oxidised?

- A. 0.907 kg B. 0.900 kg C. 1.30 kg D. 2.54 kg

(2002 NA/12)

The diagram shows the mass spectrum of a sample of naturally-occurring copper.



What is the relative atomic mass of this copper?

- A. 63.2 B. 63.5 C. 63.6 D. 64.0

(2002 NA/10)

Which pairs of compounds have the same empirical formula?

- ethane and ethene
- ethane and cyclohexane
- cyclohexane and oct-1-ene

(2002 MC/31)

A mixture of 1.0 cm³ of methane and 1.0 cm³ of ethane was sparked with an excess of oxygen. After cooling to room temperature, the residual gas was passed through aqueous potassium hydroxide.

Which compounds have the empirical formula CH_4O ?

- methanol
- ethanoic acid
- methyl methanolate

(2002 ON/31)

What is the number of molecules in 500 cm³ of oxygen under room conditions?

- A. 1.25×10^{23} B. 1.25×10^{24} C. 3.0×10^{23} D. 3.0×10^{24}

(2008 MC/12)

In the preparation of soft margarine, glycerol trioleate reacts



to saturate hydrogenated so that, on average, one of its side-chains is converted into the $\text{CH}_3(\text{CH}_2)_7\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CO}_2\text{C}$ moiety and two side-chains are converted into the $\text{CH}_3(\text{CH}_2)_7\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CH}_2\text{CO}_2\text{C}$ moiety.

How many moles of hydrogen are required to convert one mole of glycerol trioleate into the soft margarine?

- A. 4 B. 5 C. 6 D. 8

(2008 MC/22)

Use of the Data Booklet is relevant to this question.

Additional classes are about very small amounts of carbon which dissolve in 1000 cm³ of water.

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Victor M. Corman



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