

Chemistry

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(Chapter 1)(Solutions) (Intext Questions)

(Class 12)

Question 1.1:

Calculate the mass percentage of benzene (C_6H_6) and carbon tetrachloride (CCl_4) if 22 g of benzene is dissolved in 122 g of carbon tetrachloride.

Answer 1.1:

$$\begin{aligned}\text{Mass percentage of } C_6H_6 &= \frac{\text{Mass of } C_6H_6}{\text{Total mass of the solutions}} \times 100 \% \\ &= \frac{\text{Mass of } C_6H_6}{\text{Mass of } C_6H_6 + \text{Mass of } CCl_4} \times 100 \% = \frac{22}{22 + 122} \times 100 \% = 15.28 \%\end{aligned}$$

$$\begin{aligned}\text{Mass percentage of } CCl_4 &= \frac{\text{Mass of } CCl_4}{\text{Total mass of the solutions}} \times 100 \% \\ &= \frac{\text{Mass of } CCl_4}{\text{Mass of } C_6H_6 + \text{Mass of } CCl_4} \times 100 \% = \frac{122}{22 + 122} \times 100 \% = 84.72 \%\end{aligned}$$

Alternatively, Mass percentage of CCl_4 = (100 - 15.28) % = 84.72%

Question 1.2:

Calculate the mole fraction of benzene in solution containing 30% by mass in carbon tetrachloride.

Answer 1.2:

Let the total mass of the solution be 100 g and the mass of benzene be 30 g.

∴ Mass of carbon tetrachloride = (100 - 30) g = 70 g

Molar mass of benzene (C_6H_6) = (6 × 12 + 6 × 1) g mol⁻¹ = 78 g mol⁻¹

∴ Number of moles of C_6H_6 = $\frac{30}{78} \text{ mol} = 0.3846 \text{ mol}$

Molar mass of carbon tetrachloride (CCl_4) = 1 × 12 + 4 × 35.5
= 154 g mol⁻¹

∴ Number of moles of CCl_4 = $\frac{70}{154} \text{ mol} = 0.4545 \text{ mol}$

Thus, the mole fraction of C_6H_6 is given as:

$$\begin{aligned}&= \frac{\text{Number of moles of } C_6H_6}{\text{Number of moles of } C_6H_6 + \text{Number of moles of } CCl_4} \\ &= \frac{0.3846}{0.3846 + 0.4545} \\ &= 0.458\end{aligned}$$

Question 1.3:

Calculate the molarity of each of the following solutions:

(a) 30 g of $Co(NO_3)_2 \cdot 6H_2O$ in 4.3 L of solution

(b) 30 mL of 0.5 M H_2SO_4 diluted to 500 mL.

Answer 1.3:

Molarity is given by: $\text{Molarity} = \frac{\text{Moles of solute}}{\text{Volume of solution in litre}}$

(a) Molar mass of $Co(NO_3)_2 \cdot 6H_2O$ = 59 + 2 (14 + 3 × 16) + 6 × 18 = 291 g mol⁻¹

∴ Moles of $Co(NO_3)_2 \cdot 6H_2O$ = $\frac{30}{291} \text{ mol} = 0.103 \text{ mol}$

Therefore, molarity = $\frac{0.103 \text{ mol}}{4.3 \text{ L}} = 0.023 \text{ M}$

(b) Number of moles present in 1000 mL of 0.5 M H_2SO_4 = 0.5 mol

∴ Number of moles present in 30 mL of 0.5 M H_2SO_4 = $\frac{0.5 \times 30}{1000} \text{ mol} = 0.015 \text{ mol}$

Therefore, molarity = $\frac{0.015 \text{ mol}}{0.5 \text{ L}} = 0.03 \text{ M}$

8 Chemistry Sol Answers

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Sharma, Dr. B. Kapila**



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