

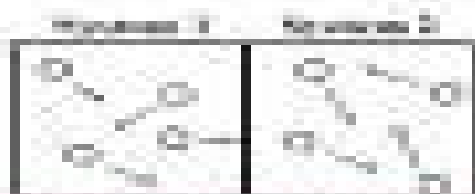
Unit 6: Thermochemistry**6.1: Heat Transfer and Thermal Equilibrium****6.2: Heat Capacity and Calorimetry****6.3: Entropy and Spontaneous Changes**

2. Use the Boltzmann probability distribution curves below to predict temperatures of the mixtures of particles.



Temperature: is a measure of the average kinetic energy of the particles in a system. The kinetic energy for particles moving with v is $\frac{1}{2}mv^2$, so the more of particles that move fast enough, the more kinetic energy we will see for a system. Therefore, the more particles move faster than the average velocity we will see the temperature's increasing (that is why the average velocity increases).

3. Sketch a curve (graph) to represent the relative distribution of the particles in the following two gases, samples of the same gas.



- a. In System 1, it is a higher or lower temperature than System 2? Justify your answer.
System 1 is at a lower temperature than System 2. Temperature is a measure of the average KE of the particles in the system. In System 1, the gas particles are spread out, so all the individual particles have the same energy, depending on their velocities KE = $\frac{1}{2}mv^2$, so all the particles have the same mass, velocity is the same magnitude but different KE, so the average in System 1 is smaller. Since particles are moving at slower velocities, the average is lower, so it is at a lower temperature.
- b. In which direction will thermal energy be transferred? Justify your answer in terms of molecular collisions.
All heat is transferred from System 2 to System 1. The particles in System 2 are moving faster than the particles in System 1 (same temperature). Therefore, the particles in System 2 are moving faster than the particles in System 1, so when they collide, the particles in System 2 will transfer some of their kinetic energy to the particles in System 1. This molecular collision will produce what is called a cooling effect on the

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