

Lab 10: Calculations with a Chemical Structure

Objectives:

1. To derive an empirical formula of a compound without using percent burning (combustion) analysis.
2. To calculate the number of moles of each of the reactants/products in the reaction.
3. To determine the limiting reagent (LR).
4. To determine the theoretical amount of product that would be produced.
5. To compare the theoretical amount to the actual amount of substance formed and calculate the percent yield.

Reaction equation:

Reaction of sodium chloride solution: 25.0 mL

Reaction of calcium chloride solution: 25.0 mL

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Reaction of 10% NaOH: 5.00 g

Procedure:

Reaction and Calculations:

1. Reaction:

$$\text{mole of sodium chloride} = 25.0 \text{ mL} \times \frac{1.0 \text{ g/mL}}{58.44 \text{ g/mol}} = 0.428 \text{ mol}$$

$$\text{mole of calcium chloride} = 25.0 \text{ mL} \times \frac{1.0 \text{ g/mL}}{110.98 \text{ g/mol}} = 0.225 \text{ mol}$$

$$\text{mole of sodium hydroxide} = 5.00 \text{ g} \times \frac{1 \text{ mol}}{40.00 \text{ g/mol}} = 0.125 \text{ mol}$$

$$\text{mole of calcium hydroxide} = 5.00 \text{ g} \times \frac{1 \text{ mol}}{74.09 \text{ g/mol}} = 0.067 \text{ mol}$$

2. Reaction:

Reaction:

$$\text{mole of } \text{Ca}^{2+} = 0.225 \text{ mol} \times \frac{1 \text{ mol}}{1 \text{ mol}} = 0.225 \text{ mol}$$

$$\text{mole of } \text{Ca}^{2+} = 0.225 \text{ mol} \times \frac{1 \text{ mol}}{1 \text{ mol}} = 0.225 \text{ mol}$$

$$\text{mole of } \text{Ca}^{2+} = 0.225 \text{ mol} \times \frac{1 \text{ mol}}{1 \text{ mol}} = 0.225 \text{ mol}$$

Answers For Lab 6b Heath Chemistry

Arturo Cuomo



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