

Oxidation-Reduction Worksheet

For each reaction below, identify the atom oxidized, the atom reduced, the oxidizing agent, the reducing agent, the oxidation half reaction, the reduction half reaction, and then balance the equation by the method of oxidation-reduction showing all electrons transfers.

1. $\overset{0}{\text{Mg}} + 2\overset{+1}{\text{H}}\overset{-1}{\text{Cl}} \rightarrow \overset{+2}{\text{Mg}}\overset{-2}{\text{Cl}}_2 + \overset{0}{\text{H}}_2$ $\overset{0}{\text{Mg}} \xrightarrow{-2e^-} \overset{+2}{\text{Mg}}$ $2\overset{+1}{\text{H}} \xrightarrow{+2e^-} \overset{0}{\text{H}}_2$
2. $2\overset{0}{\text{Fe}} + 3\overset{+3}{\text{V}}_2\overset{-2}{\text{O}}_3 \rightarrow \overset{+3}{\text{Fe}}_2\overset{-2}{\text{O}}_3 + 6\overset{+2}{\text{V}}\overset{-2}{\text{O}}$ $2\overset{0}{\text{Fe}} \xrightarrow{-6e^-} \overset{+3}{\text{Fe}}_2$ $3\overset{+3}{\text{V}}_2 \xrightarrow{+6e^-} 6\overset{+2}{\text{V}}$
3. $\overset{+7}{\text{K}}\overset{-2}{\text{Mn}}\overset{-1}{\text{O}}_4 + \overset{+5}{\text{K}}\overset{-2}{\text{N}}\overset{-1}{\text{O}}_2 + 2\overset{+2}{\text{H}}_2\overset{-2}{\text{S}}\overset{-1}{\text{O}}_4 \rightarrow \overset{+2}{\text{Mn}}\overset{-2}{\text{S}}\overset{-1}{\text{O}}_4 + 2\overset{+2}{\text{H}}_2\overset{-2}{\text{O}} + \overset{+5}{\text{K}}\overset{-2}{\text{N}}\overset{-1}{\text{O}}_3 + \overset{+6}{\text{K}}_2\overset{-2}{\text{S}}\overset{-1}{\text{O}}_4$
4. $\text{K}_2\text{Cr}_2\text{O}_7 + \text{SnCl}_2 + \text{HCl} \rightarrow \text{CrCl}_3 + \text{SnCl}_4 + \text{KCl} + \text{H}_2\text{O}$
5. $\text{KMnO}_4 + \text{NaCl} + \text{H}_2\text{SO}_4 \rightarrow \text{Cl}_2 + \text{K}_2\text{SO}_4 + \text{MnSO}_4 + \text{H}_2\text{O} + \text{Na}_2\text{SO}_4$
6. $\text{K}_2\text{Cr}_2\text{O}_7 + \text{H}_2\text{O} + \text{S} \rightarrow \text{SO}_2 + \text{KOH} + \text{Cr}_2\text{O}_3$
7. $\text{KClO}_3 + \text{C}_{12}\text{H}_{22}\text{O}_{11} \rightarrow \text{KCl} + \text{H}_2\text{O} + \text{CO}_2$
8. $\text{H}_2\text{C}_2\text{O}_4 + \text{K}_2\text{MnO}_4 \rightarrow \text{CO}_2 + \text{K}_2\text{O} + \text{Mn}_2\text{O}_3 + \text{H}_2\text{O}$
9. $\text{Mn}(\text{NO}_3)_2 + \text{NaBiO}_3 + \text{HNO}_3 \rightarrow \text{HMnO}_4 + \text{Bi}(\text{NO}_3)_3 + \text{NaNO}_3 + \text{H}_2\text{O}$
10. $\text{H}_2\text{C}_2\text{O}_4 + \text{KMnO}_4 \rightarrow \text{CO}_2 + \text{K}_2\text{O} + \text{Mn}_2\text{O}_3 + \text{H}_2\text{O}$

Analyzing Oxidation Reduction Reactions Answer Key

WJ Hussar



Analyzing Oxidation Reduction Reactions Answer Key:

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In a fast-paced world fueled by information and interconnectivity, the spellbinding force of linguistics has acquired newfound prominence. Its capacity to evoke emotions, stimulate contemplation, and stimulate metamorphosis is really astonishing. Within the pages of "**Analyzing Oxidation Reduction Reactions Answer Key**," an enthralling opus penned by a very acclaimed wordsmith, readers attempt an immersive expedition to unravel the intricate significance of language and its indelible imprint on our lives. Throughout this assessment, we shall delve to the book is central motifs, appraise its distinctive narrative style, and gauge its overarching influence on the minds of its readers.

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