

Задача 11.
Слабые и сред. кислоты

1) Дано:

$$C_{HNO_2} = 0,01 \text{ м}$$

$$K_a = 7,9 \cdot 10^{-4}$$

$$pH = ?$$

$$pH = -\lg [H^+] = -\lg 2,8 \cdot 10^{-5} = 4,55$$

Решение:

Для слабой кислоты:

$$[H^+] = \sqrt{K_a \cdot C_{HNO_2}}$$

$$= \sqrt{7,9 \cdot 10^{-4} \cdot 10^{-2}} = 2,8 \cdot 10^{-5} \text{ моль/л}$$

2) Дано:

$$K_{Pr} = 6,3 \cdot 10^{-5}$$

$$pH = ?$$

$$S = \sqrt[3]{\frac{K_{Pr}}{4}} = \sqrt[3]{\frac{6,3 \cdot 10^{-5}}{4}} = 1,16 \cdot 10^{-2} \text{ моль/л}$$

$$[OH^-] = \frac{S}{2} = \frac{1,16 \cdot 10^{-2}}{2} = 5,8 \cdot 10^{-3} \text{ моль/л}$$

$$[H^+] = \frac{10^{-14}}{[OH^-]} = \frac{10^{-14}}{5,8 \cdot 10^{-3}} = 1,72 \cdot 10^{-12}$$

$$pH = -\lg [H^+] = -\lg 1,72 \cdot 10^{-12} = 11,76$$

Решение:

$$Ca(OH)_2 \rightleftharpoons Ca^{2+} + 2OH^-$$

$$K_{Pr} = [Ca^{2+}][OH^-]^2$$

$$= 6 \cdot (2S)^2 = 12S^3$$

3) Дано:

$$V_{HCl} = 10 \text{ мл}$$

$$C_{HCl} = 0,1 \text{ м}$$

$$V_{NaClO_2} = 10 \text{ мл}$$

$$C_{NaClO_2} = 0,1 \text{ м}$$

$$K_{Pr} = 4,6 \cdot 10^{-5}$$

$$PrCl_2 \rightleftharpoons Pr^{2+} + 2Cl^-$$

Решение:

полн. соль:

$$C_{HCl} = \frac{0,1 \cdot 10}{20} = 0,05 \text{ моль/л}$$

$$C_{NaClO_2} = \frac{0,1 \cdot 10}{20} = 0,05 \text{ моль/л}$$

полн. гидролизующая соль:

$$C_{Cl^-} = C_{HCl} = 0,05 \text{ моль/л}$$

$$C_{Pr^{2+}} = C_{NaClO_2} = 0,05 \text{ моль/л}$$

$$[Pr^{2+}][Cl^-]^2 = 4,6 \cdot 10^{-5} = 10 \cdot x^3$$

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R Sanford



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1. Understanding the eBook Answers For Apex Physics Semester 2
 - The Rise of Digital Reading Answers For Apex Physics Semester 2
 - Advantages of eBooks Over Traditional Books
2. Identifying Answers For Apex Physics Semester 2
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers For Apex Physics Semester 2
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answers For Apex Physics Semester 2
 - Personalized Recommendations
 - Answers For Apex Physics Semester 2 User Reviews and Ratings
 - Answers For Apex Physics Semester 2 and Bestseller Lists
5. Accessing Answers For Apex Physics Semester 2 Free and Paid eBooks
 - Answers For Apex Physics Semester 2 Public Domain eBooks
 - Answers For Apex Physics Semester 2 eBook Subscription Services
 - Answers For Apex Physics Semester 2 Budget-Friendly Options

6. Navigating Answers For Apex Physics Semester 2 eBook Formats
 - ePub, PDF, MOBI, and More
 - Answers For Apex Physics Semester 2 Compatibility with Devices
 - Answers For Apex Physics Semester 2 Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers For Apex Physics Semester 2
 - Highlighting and Note-Taking Answers For Apex Physics Semester 2
 - Interactive Elements Answers For Apex Physics Semester 2
8. Staying Engaged with Answers For Apex Physics Semester 2
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers For Apex Physics Semester 2
9. Balancing eBooks and Physical Books Answers For Apex Physics Semester 2
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answers For Apex Physics Semester 2
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Answers For Apex Physics Semester 2
 - Setting Reading Goals Answers For Apex Physics Semester 2
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Answers For Apex Physics Semester 2
 - Fact-Checking eBook Content of Answers For Apex Physics Semester 2
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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