

Analysis of Alum $\text{KAl}(\text{SO}_4)_2 \cdot 12\text{H}_2\text{O}$

After a compound has been synthesized, tests should be carried out to verify the compound formed is indeed the compound desired. There are a number of various tests that can be performed to verify that the compound is the one desired. In a previous experiment, alum crystals were prepared. In this experiment we will do several tests to determine if the crystals are really alum.

The first and simplest test is to find the melting point of the compound and compare it to the published value for alum. A small quantity of alum is powdered and placed on a melting point apparatus. The crystals are heated on the melting point apparatus on a hot plate, and the temperature at which they melt is recorded and compared to the reported values.

The second test that we can do is to determine the amount of water of hydration present in the alum crystals. Some of the alum is placed in a crucible and weighed. The crucible is heated until all of the water of hydration is driven off. The crucible is then cooled and its mass measured. From the mass of the dry crystals and the mass of the water lost, the ratio of moles water to moles $\text{KAl}(\text{SO}_4)_2$ can be calculated and then compared to the correct formula values.

The third test is a chemical test to determine the percent of sulfate in the compound prepared. A weighed quantity of alum is dissolved in distilled water. An excess of barium ions is added to the solution to precipitate all of the sulfate in the compound as barium sulfate. The precipitated barium sulfate is filtered off, dried, and its mass determined. From the mass of the barium sulfate and the initial mass of the alum, the percent sulfate can be calculated and compared with the theoretical percent found from the formula.

Chemicals

Parts 1, 2, and 3

Alum crystals prepared by you in the Lab
Known Alum crystals

Part 3

Barium nitrate, $\text{Ba}(\text{NO}_3)_2$, 0.2M

Equipment

Part 1

Capillary Tube	Thermometer
Melting point apparatus	Mortar and pestle
Hot plate	

Part 2

Crucible and cover	Ring stand, ring and Bunsen burner
Clay triangle	

Analysis Of Alum Lab Answers Sulfate

Sheridan Muspratt



Analysis Of Alum Lab Answers Sulfate:

Laboratory Study of Chemical Coagulation as a Means of Treatment for Dredged Material Chun-Ching Wang, Junqing Wang, Kenneth Y. Chen, 1977 This report discusses trace metals chlorinated hydrocarbons and biostimulants which are found to be enriched in fine particulates of contaminated sediments When different types of sediments are resuspended it has been found that the extent of redissolution is very small However contaminants attached to the particulates can be transported along the waterways downstream from the disturbed areas In confined areas proper treatment and rapid removal of those suspended particulates from the water column will not only lessen the oxygen demand but will also reduce the input of contaminants into the receiving water **Chemistry, Theoretical, Practical, and Analytical** Sheridan Muspratt, 1860 *Report* North Carolina Agricultural Experiment Station, 1899 Vol for 1889 contains also Bulletin no 67 a Technical bulletin no 1 and Annual report of the Meteorological Division for the North Carolina Agricultural Experiment Station constituting the N C State Weather Service for 1889 vols for 1894 contain also Bulletin no 94 also Press bulletin also 1916 17 1918 19 Technical bulletins no 11 16 **Engineering and Mining Journal**, 1955

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