Limiting Reagent & Percent Yield Practice Worksheet

- When copper (II) chloride reacts with sodium nitrate, copper (II) nitrate and sodium chloride are formed.
 - a. Write the balanced equation for the reaction given above:

+ NaNO₁ → Cu(NO₁)₂ + ChrCile CuCl₂ + 2 NaNO₃ → Cu(NO₃)₂ + 2 NaCl

 If 15 grams of copper (II) chloride react with 20 grams of sodium nitrate, how much sodium chloride can be formed?

1 mol CuCl₂ x 2 mol NaCl x 58.443 g NaCl = 13 g NaCl 15 g CuClax 134.362g CuCl: 1 mol CuCl: 1 mol NaCl

20 g NaNO₃ x 1 mol NaNO₃ x 2 mol NaCl x 58.443 g NaCl = 13.8 g NaCl 84,997g NaNOs 2 mol NaNOs 1 mol NaCl

Only 13 g can be formed

- c. What is the limiting reagent for the reaction in #2? copper (II) chloride
- d. How many grams of copper (II) nitrate is formed?

 $15g CuCl_2 \times 1 \mod CuCl_2 \times 1 \mod Cu(NO_3)_2 \times 187.467 g Cu(NO_3)_2 = 20.9 g Cu(NO_3)_2$ 134,452g CuCl₂ 1 mol CuCl₂ 1 mol Cu(NO₃)₂

e. How much of the excess reasent is left over in this reaction?

15g CuCl₂ x 1 mol CuCl₂ x 2 mol NaNO₃ x 84.997g NaNO₃ = 19 g NaNO₃ 134.452g CuCl₂ 1 mol CuCl₂ 1 mol NaNO₃

20 g NaNOs - 19 g NaNOs = 1 g NaNOs

f. If 11.3 grams of sodium chloride are formed in the reaction described in problem #2, what is Percent yield = $\frac{\text{actual yield}}{\text{theoretical yield}} \times 100 \text{ percent}$ the percent yield of this reaction?

formed.

11.3/13.0 x 100% = 86.9%

- 2. When lead (II) nitrate reacts with sodium iodide, sodium nitrate and lead (II) iodide are
 - a. Balance the following equation:

Pb(NO₃)_{2 (and} + NaI (and → PbI_{2 (a)} + NaNO_{3 (and} Pb(NO₃)_{2 tant} + 2 NaI _{tant} → PbI_{2 tot} + 2 NaNO_{3 tant}

 If I start with 25.0 grams of lead (II) nitrate and 15.0 grams of sodium iodide, how many grams of sodium nitrate can be formed?

25.0 g Pb(NO₃)₂ x 1 mol Pb(NO₃)₂ x 2 mol NaNO₃ x 84.997g NaNO₃ = 12.8 g NaNO₃ 331.214 g Pb(NO₃)₂ 1 mol Pb(NO₃)₂ 1 mol NaNO₃

15.0 g Nal x 1 mol Nal x 2 mol NaNO₃ x 84.997g NaNO₃ = 8.51 g NaNO₃ 149,89g Nat 2 mol NaI 1 mol NaNOs

8.51 g NaNO3 can be made

Answers Limiting Reactants And Percent Yield

Howard Stephen Stoker

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Table of Contents Answers Limiting Reactants And Percent Yield

- 1. Understanding the eBook Answers Limiting Reactants And Percent Yield
 - The Rise of Digital Reading Answers Limiting Reactants And Percent Yield
 - Advantages of eBooks Over Traditional Books
- 2. Identifying Answers Limiting Reactants And Percent Yield
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answers Limiting Reactants And Percent Yield
 - User-Friendly Interface
- 4. Exploring eBook Recommendations from Answers Limiting Reactants And Percent Yield

- Personalized Recommendations
- Answers Limiting Reactants And Percent Yield User Reviews and Ratings
- Answers Limiting Reactants And Percent Yield and Bestseller Lists
- 5. Accessing Answers Limiting Reactants And Percent Yield Free and Paid eBooks
 - Answers Limiting Reactants And Percent Yield Public Domain eBooks
 - Answers Limiting Reactants And Percent Yield eBook Subscription Services
 - Answers Limiting Reactants And Percent Yield Budget-Friendly Options
- 6. Navigating Answers Limiting Reactants And Percent Yield eBook Formats
 - o ePub, PDF, MOBI, and More
 - Answers Limiting Reactants And Percent Yield Compatibility with Devices
 - Answers Limiting Reactants And Percent Yield Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answers Limiting Reactants And Percent Yield
 - Highlighting and Note-Taking Answers Limiting Reactants And Percent Yield
 - Interactive Elements Answers Limiting Reactants And Percent Yield
- 8. Staying Engaged with Answers Limiting Reactants And Percent Yield
 - o Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answers Limiting Reactants And Percent Yield
- 9. Balancing eBooks and Physical Books Answers Limiting Reactants And Percent Yield
 - Benefits of a Digital Library
 - o Creating a Diverse Reading Collection Answers Limiting Reactants And Percent Yield
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answers Limiting Reactants And Percent Yield
 - Setting Reading Goals Answers Limiting Reactants And Percent Yield
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answers Limiting Reactants And Percent Yield

- Fact-Checking eBook Content of Answers Limiting Reactants And Percent Yield
- Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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