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Student Exploration: Diffusion

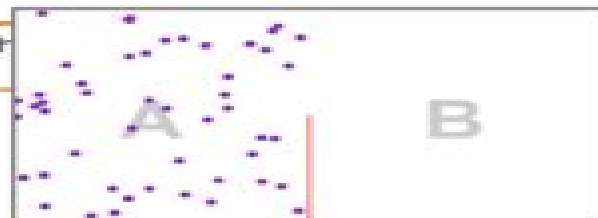
Directions: Follow the instructions to go through the simulation. Respond to the questions and prompts in the orange boxes.

Vocabulary: absolute zero, controlled experiment, diffusion, dynamic equilibrium, Kelvin scale, kinetic energy

Prior Knowledge Question (Do this BEFORE using the Gizmo.)

Have you ever smelled microwave popcorn? The oddly enticing scent can fill a whole house. How do you think the smell of popcorn spreads through the air?

This is through the process of diffusion, the gases spread throughout the chamber from high concentration to low concentration.



Gizmo Warm-up

Smells are carried by tiny particles that move through the air. The *Diffusion Gizmo* shows gas particles in a chamber that is divided into two regions by a partial wall. Click **Play** () and observe.

1. Describe the motion of the gas particles.

The gas particles are fast-moving and constantly in motion, bouncing off the walls. In the beginning, many particles move to side B as it has a lower concentration than side A.

2. Over time, what is happening?

Over time, the concentrations of particles even out on both sides (become isotonic), due to the movement of particles from a high concentration to a low concentration.

This process, in which particles move from an area of high concentration (region A) to an area of low concentration (region B), is called **diffusion**.

3. Select the BAR CHART tab, and observe the chart for a few minutes. After the first 30 seconds or so, how much do the numbers of particles in each region change?

The numbers of particles in each region do not change drastically after 30 seconds of net movement. For example, in region B at 0 seconds, the number of particles was 0 and only increased to 10 at 30 seconds.

When the numbers don't change much, the particles are said to be in **dynamic equilibrium**.

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