

Assignment #2 – Compound Names and Formulas
Multi-valent ions only

A. Write the correct formula for the following compounds, all of which have been named using the modern Roman Numeral Method. The ionic charge is given after the first element (metallic) in Roman Numerals.

1. copper (II) oxide CuO
2. mercury (I) oxide Hg₂O
3. gold (III) chloride AuCl₃
4. thallium (III) bromide TlBr₃
5. bismuth (V) oxide Bi₂O₅
6. terbium (III) oxide Tb₂O₃
7. uranium (VI) oxide UO₃
8. protactinium (IV) bromide PaBr₄
9. cerium (III) oxide Ce₂O₃
10. arsenic (V) sulphide As₂S₅

11. manganese (III) oxide Mn₂O₃
12. vanadium (II) bromide VBr₂
13. niobium(V) oxide Nb₂O₅
14. titanium (III) oxide Ti₂O₃
15. titanium (III) nitride TiN
16. iron (II) oxide FeO
17. cobalt (II) phosphide Co₃P₂
18. tin (II) oxide SnO
19. thulium (II) bromide TmBr₂
20. copper (II) bromide CuBr₂

B. Name the compound using the Roman Numeral Method and balance the ionic charges.

1. SnCl₄ Tin (IV) Chloride
2. BiBr₅ Bismuth (V) Bromide
3. PoO₂ Polonium (IV) Oxide
4. PbI₂ Lead (II) Iodide
5. HgO Mercury (II) Oxide
6. HgCl Mercury (I) Chloride
7. Au₂O₃ Gold (III) Oxide
8. FeCl₂ Iron (II) Chloride

9. PdF₄ Palladium (IV) Fluoride
10. Os₂O₃ Osmium (III) Oxide
11. MoBr₂ Molybdenum (II) Bromide
12. VCl₅ Vanadium (V) Chloride
13. Mn₂O₃ Manganese (III) Oxide
14. CoO Cobalt (II) Oxide
15. Np₂O₃ Neptunium (III) Oxide
16. V₂O₅ Vanadium (V) Oxide

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RD Boyd



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Table of Contents Answer Key Chemical Formulas Chemical Formulas

1. Understanding the eBook Answer Key Chemical Formulas Chemical Formulas
 - The Rise of Digital Reading Answer Key Chemical Formulas Chemical Formulas
 - Advantages of eBooks Over Traditional Books
2. Identifying Answer Key Chemical Formulas Chemical Formulas
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Answer Key Chemical Formulas Chemical Formulas
 - User-Friendly Interface
4. Exploring eBook Recommendations from Answer Key Chemical Formulas Chemical Formulas
 - Personalized Recommendations
 - Answer Key Chemical Formulas Chemical Formulas User Reviews and Ratings
 - Answer Key Chemical Formulas Chemical Formulas and Bestseller Lists
5. Accessing Answer Key Chemical Formulas Chemical Formulas Free and Paid eBooks
 - Answer Key Chemical Formulas Chemical Formulas Public Domain eBooks
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- Answer Key Chemical Formulas Chemical Formulas Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Answer Key Chemical Formulas Chemical Formulas
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 - Interactive Elements Answer Key Chemical Formulas Chemical Formulas
- 8. Staying Engaged with Answer Key Chemical Formulas Chemical Formulas
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Answer Key Chemical Formulas Chemical Formulas
- 9. Balancing eBooks and Physical Books Answer Key Chemical Formulas Chemical Formulas
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Answer Key Chemical Formulas Chemical Formulas
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Answer Key Chemical Formulas Chemical Formulas
 - Setting Reading Goals Answer Key Chemical Formulas Chemical Formulas
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Answer Key Chemical Formulas Chemical Formulas
 - Fact-Checking eBook Content of Answer Key Chemical Formulas Chemical Formulas
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
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