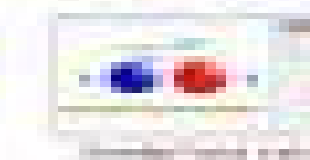
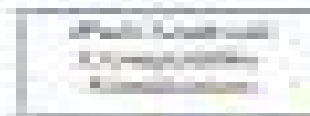


Experiment 1: The Effect of Temperature on the Rate of Reaction

Name: _____

Background Information

Every science experiment is designed to test a hypothesis. In this experiment, the hypothesis is that the rate of reaction will increase as the temperature of the reactants increases. The rate of reaction is the change in concentration of the products per unit time. The rate of reaction can be measured by the change in the amount of the products formed over a given time interval. The rate of reaction can also be measured by the change in the amount of the reactants consumed over a given time interval. The rate of reaction can be measured by the change in the amount of the products formed over a given time interval. The rate of reaction can also be measured by the change in the amount of the reactants consumed over a given time interval.



Procedure

1. Prepare a solution of sodium carbonate in water. The concentration of the solution should be 1.0 g/dm³.
2. Measure 10 cm³ of the solution into a conical flask.

3. Measure 10 cm³ of dilute hydrochloric acid into a measuring cylinder.



4. Add the acid to the flask and start the stopwatch. Record the time taken for the reaction to occur.

5. Repeat the experiment at different temperatures.

1. Prepare a solution of sodium carbonate in water. The concentration of the solution should be 1.0 g/dm³. Measure 10 cm³ of the solution into a conical flask.
2. Measure 10 cm³ of dilute hydrochloric acid into a measuring cylinder.
3. Add the acid to the flask and start the stopwatch. Record the time taken for the reaction to occur.
4. Repeat the experiment at different temperatures.

Volume of acid / cm ³	Volume of alkali / cm ³	Temperature / °C	Time taken / s	Rate of reaction / g/dm ³ s
10	10	20	120	0.0083
10	10	30	80	0.0125
10	10	40	60	0.0167
10	10	50	40	0.0250
10	10	60	30	0.0333
10	10	70	20	0.0500
10	10	80	15	0.0667
10	10	90	10	0.1000
10	10	100	5	0.2000

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Lei Huang



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