

Qualitative Analysis and Chemical Bonding

Concepts:

- Chemical bonding
- Covalent bonding
- Ionic bonding
- Metallic bonding
- Intermolecular bonding
- Physical & chemical properties
- Electronegativity

Introduction

- Types of atoms are not equal in chemical bonds: **electronegativity**
- In order to understand how all substances behave we need to know how they are bonded, strong
- In order to know a substance behaves like ionic, strong, must be able to separate its ions completely

Electronegativity

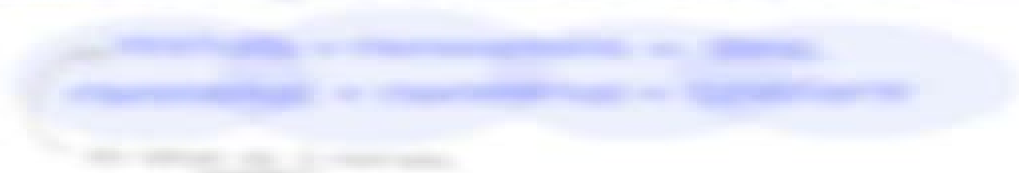
- Electronegativity: tendency of an atom to attract electrons towards itself in a covalent compound
- More electronegative atom is more able to attract electrons in a bond, pulls electrons more strongly, more able
- Electronegativity increases with the strength of bonds
- Periodic table electronegativity decreases **downward**
- From electronegative fluorine (strongly attracts electrons) to weakly attracts an **alkali metal**



Covalent bonding

Two electronegative atoms sharing their outer, or valence, electrons to form a bond.

Electron pair is a group of atoms that together are neutral overall.



Electron pair bond is stronger than ionic bond. Ionic bond is weaker, but still strong. Ionic bond is a transfer of electrons, while covalent bond is sharing of electrons.

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Answer Key To Qualitative Analysis And Chemical Bonding

Gernot Frenking, Sason Shaik



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