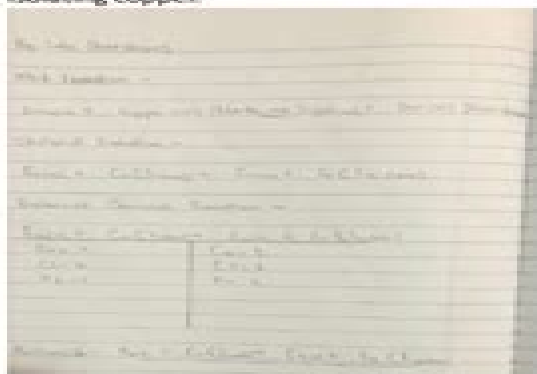


Limiting Reagents and % Yield

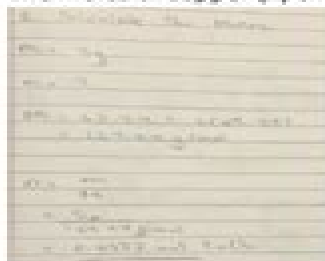
The percentage yield of a chemical reaction is determined by numerous factors: the nature of the reaction itself, the conditions under which the reaction was carried out and the nature of the reactants used. In this investigation, you will determine the limiting/excess reagents and the percentage yield of the reaction between iron and a solution of copper (II) chloride.

Prelab Questions:

- Write the balanced chemical equation for the reaction between iron and copper (II) chloride. Remember to use your activity series to check the products.
 - The balanced equation is $\text{Fe}(s) + \text{CuCl}_2(aq) \rightarrow \text{Cu}(s) + \text{FeCl}_2(aq)$. The reaction is a single displacement isolating copper.



- Calculate the moles of copper (II) chloride in the solution if you use 5 g of copper (II) chloride.
 - The moles of copper (II) chloride in the solution, if I use 5g of copper (II) chloride, is 0.037 mol CuCl₂.



- Determine the limiting reagent if you use 1 g of iron.
 - The limiting reagent, if I use 1g of iron, is Fe(s). It is 0.018mol because when comparing the numbers by the

Limiting Reagents and % Yield



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