

(C) Researcher says about the evidence, and then presents further evidence.

(D) First presents the evidence, and then provides a way for the audience to see the evidence. (1)

2222. Which of the following statements is true about evidence after you've answered?

(A) They are all correct for the developing paragraph.

(B) They are all correct for all developing paragraphs.

(C) Their position and alignment are important.

(D) They are not at all for telling. (1)

2223. A reader has decided: Which of the following is a common sign of a bad start to a paragraph?

(A) None

(B) Evidence

(C) None

(D) Telling (1)

2224. When reading a sentence, frequent signaling on the right side of the sentence is

(A) Telling

Answers Mosby 6th Edition Work

Sandy Fritz, Luke Allen Fritz



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web the four different types of orbitals s p d and f have different shapes and one orbital can hold a maximum of two electrons the p d and f orbitals have different sublevels thus can hold more electrons as stated the electron configuration of each element is unique to its position on the periodic table

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web explanation the first two electrons fit into the 1s orbital 1s 2 then next two into the 2s orbital 2s 2 the 2 p orbital has three slots which can take 2 electrons each a total of 6 which gives 2p 6 the final two electrons fit in the 3s orbital 3s 2

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web using the periodic table to determine the electron configurations of atoms is key but also keep in mind that there are certain rules to follow when assigning electrons to different orbitals the periodic table is an incredibly helpful tool in writing electron configurations

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web explain your answer the first electron to leave the atom would be one in the highest energy level which is the fourth energy level electrons in the highest energy level are the least attracted to the nucleus because they are the most distant 25

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web electron configurations worksheet with answers electron configuration notation shows the arrangement of electrons around the nucleus of an atom helps chemist understanding how elements form chemical bonds can be written using the period table or an electron configuration chart

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web the following electron configurations belong to which elements 21 1s22s22p63s1 sodium 22

1s22s22p63s23p64s23d104p65s24d6 ruthenium 23 kr 5s24d10 cadmium 24 xe 6s24f145d106p2 lead 25 rn 7s25f146d4

seaborgium determine if the following electron configurations are correct 26 1s22s22p63s23p64s24d104p65s1 no it should be 3d10

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web how many protons neutrons and electrons are in atoms of these isotopes write the complete electron configuration for each isotope answer co has 27 protons 27 electrons and 33 neutrons 1s 2 2s 2 2p 6 3s 2 3p 6 4s 2 3d 7 i has 53 protons 53 electrons and 78 neutrons 1s 2 2s 2 2p 6 3s 2 3p 6 3d 10 4s 2 4p 6 4d 10 5s 2 5p 5

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web select the electron configuration tab on show number of electrons introduction beyond argon it is a bit tricky to determine which subshell gets filled next there are several rules that scientists use to determine the electron configurations of larger atoms question how are the electron configurations of elements beyond argon determined